Question FBQ1 : \_\_\_ is a group 2 metal that is not counted as an alkaline earth metal since its oxide is not alkaline
Answer: Beryllium

Question FBQ2 : Group 2A elements are called alkaline earth metals because they react with oxygen to form \_\_\_\_\_\_\_
Answer: Oxides

Question FBQ3 : Alkaline earth metals have \_\_\_\_\_\_ valence electrons
Answer: Two

Question FBQ4 : Alkaline earth metals from ions with a \_\_\_\_\_\_ charge
Answer: +2

Question FBQ5 : The polarizing power of a cation is \_\_\_\_\_\_\_\_ proportional to its size
Answer: inversely

Question FBQ6 : Group 1 metals are known as \_\_\_\_\_\_
Answer: alkaline metals

Question FBQ7 : Group 2 elements are also known as \_\_\_\_\_\_\_ earth metals
Answer: alkaline

Question FBQ8 : John Newlands, an English chemist reported his law of \_\_\_\_\_\_\_\_\_
Answer: Octaves

Question FBQ9 : If elements are arranged sequentially in order of increasing \_\_\_\_\_\_\_\_\_\_, a periodic repetition, that is, periodicity in properties is observed
Answer: Atomic weight

Question FBQ10 : The modern periodic law states that the properties of elements are \_\_\_\_\_\_ functions of their atomic numbers
Answer: periodic

Question FBQ11 : The systematic name for the element with atomic number of 105 is called \_\_\_\_\_\_
Answer: Unnilpentium

Question FBQ13 : According to \_\_\_\_\_\_\_\_\_, no two electrons in the same atom can have the same value of n, l and mi.
Answer: Exclusion principle

Question FBQ14 : According to \_\_\_\_\_ rule, every orbital in a subshell is singly occupied with one electron before any one orbital is doubly occupied, and all electrons in singly occupied orbitals have the same spin
Answer: Hunds

Question FBQ15 : Elements in the same period have same number of \_\_\_\_\_\_\_
Answer: Shell

Question FBQ16 : Elements in the same group have the same number of \_\_\_\_\_\_\_\_
Answer: Valence electrons

Question FBQ17 : The electronic configuration of an atom with atomic number 9 is given as \_\_\_\_\_\_
Answer: 1s2 2s2 2p5

Question FBQ18 : The ionic radii of metallic cations are \_\_\_\_ than the atomic radii
Answer: Smaller

Question FBQ19 : The ionic radii of non-metallic anion are \_\_\_\_ than the atomic radii
Answer: Bigger

Question FBQ20 : Electron affinity \_\_\_\_\_\_\_ across the period
Answer: Increases

Question FBQ21 : Complete the following reaction Na2C2 + 2H2O → 2NaOH + \_\_\_
Answer: C2H2

Question FBQ22 : When Li2CO3 is decomposed under heat it gives \_\_\_\_ and CO2
Answer: Li2O

Question FBQ23 : KO2 is an example of a \_\_\_\_ oxide
Answer: Super

Question FBQ24 : When a metal is surrounded by solvent molecules, the phenomenon is called \_\_\_\_\_\_
Answer: Solvation

Question FBQ25 : Electrons associated with the solvent are known as \_\_\_\_\_\_ electrons
Answer: Solvated

Question FBQ26 : \_\_\_\_\_\_\_\_ is an alkaline earth metals useful in the formation of bones and teeth
Answer: Magnesium

Question FBQ27 : Complete the following reaction; Ca(OH)2 + CO2 → \_\_\_\_ + H2O
Answer: CaCO3

Question FBQ28 : All group 2 elements except \_\_\_ form hydrides
Answer: Beryllium

Question FBQ29 : Complex formation is favoured by \_\_\_\_, highly charged cations with suitable empty orbitals of approximately the right energy with which the ligand orbitals can combine
Answer: Small

Question FBQ30 : In chlorophyll \_\_\_\_ is coordinated to four nitrogen atoms in the heterocyclic porphyrin ring system
Answer: Magnesium

Question FBQ31 : The hydrated ionic radii of alkali metal ions \_\_\_\_\_ down the group
Answer: Decreases

Question FBQ32 : The tendency to form complexes \_\_\_\_\_\_ with increasing atomic number in alkali earth metals
Answer: Decreases

Question FBQ33 : The density of alkali earth metals shrinks from beryllium to calcium, but \_\_\_\_ considerably thereafter up to radium
Answer: Increases

Question FBQ12 : The systematic name for the element with atomic number 102 is called \_\_\_\_\_\_

Answer: Unilbium

Question FBQ34 : \_\_\_\_ is a metallic cation that is isoelectronic to Al3+
Answer: Mg2+

Question FBQ35 : Potassium (19K) has an effective Z value of \_\_\_
Answer: 2.20

Question MCQ1 : Alkaline earth metals form ions with a \_\_\_\_\_\_ charge
Answer: +2

Question MCQ2 : Which element is not an alkaline earth metals?
Answer: zirconium

Question MCQ3 : The classification of certain groups of three elements into TRIADS was done by?
Answer: Dobereiner

Question MCQ4 : Classification of elements according to increasing atomic weight on a line which spiraled around a cylinder from the bottom to top was done by ?
Answer: De chancourtois

Question MCQ5 : The Law of octaves was put forward by?
Answer: Newlands

Question MCQ6 : Arranging elements in increasing atomic weight leading to a periodicity in their properties was done by
Answer: Mendeleev

Question MCQ7 : Both elements of the 1st period contains valence electrons in
Answer: K shell

Question MCQ8 : In periodic table, helium is placed at
Answer: top right corner

Question MCQ9 : On the basis of electronic configuration group and period an element with atomic number of 5 is
Answer: 2 and IIIA

Question MCQ10 : Chemical properties depends upon
Answer: valence shell electronic configuration

Question MCQ11 : Nobel gases are present in
Answer: VIIIA group

Question MCQ12 : Energy required to remove an electron from the outermost shell is called
Answer: ionization energy

Question MCQ13 : As we go from top to bottom in a group shielding effect
Answer: Increases

Question MCQ14 : Group A elements are called
Answer: transition elements

Question MCQ15 : Period number of element indicates
Answer: value of valence shell

Question MCQ16 : Ability of atom to attract electrons towards itself is called
Answer: Electronegativity

Question MCQ17 : Average distance between nucleus and outer shell is called
Answer: atomic size

Question MCQ18 : Groups containing alkaline earth metals are
Answer: IIA

Question MCQ19 : Physical properties depends on the
Answer: size of atom

Question MCQ20 : Elements that lie in same column have
Answer: similar properties

Question MCQ21 : As we go from left to right across period, electron affinity
Answer: Increases

Question MCQ22 : Elements are arranged in order of
Answer: increasing atomic number

Question MCQ23 : Decrease in force of attraction between valence electrons and nucleus by innver electrons is called
Answer: shielding effect

Question MCQ24 : 14 elements after actinium is called
Answer: Actinides

Question MCQ25 : an element that has an atomic number of 15 with which of the following elements will it show similar chemical properties
Answer: N (7)

Question MCQ26 : The group number and period number respectively of an element with atomic number 8 is
Answer: 6, 2

Question MCQ27 : Identify the wrong sequence of elements in a group
Answer: Cu, Au, Ag

Question MCQ28 : An element with atomic number \_\_\_\_\_\_ will form a basic oxide
Answer: 11

Question MCQ29 : What principle/rule is violated in the configuration 1s2 2s3
Answer: Pauli's exclusion principle

Question MCQ30 : What principle/ rule is violated in the configuration 1s2 2s2 2px2 2py1
Answer: Hund's rule

Question MCQ31 : What principle/ rule is violated in the electronic configuration 1s2 2px2
Answer: Aufbau's principle

Question MCQ32 : What type of bonding is found in Ammonia molecule
Answer: Covalent

Question MCQ33 : The polarizing power of a cation is \_\_\_\_\_\_\_\_\_ proportional to its size
Answer: Inversely

Question MCQ34 : Na3AlF6 is known as
Answer: Cryolite

Question MCQ35 : When a metal is surrounded by solvent molecules, the phenomenon is called
Answer: Solvation