

eExam Question Bank

Coursecode:

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<input type="checkbox"/>	Question Type	Question	A	B	C	D	Answer	Remark
<input type="checkbox"/>	FBQ	Hydride ion is unusually <input type="text"/> with respect to size.	Large	Large				<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	The stability of alkali salts depend on <input type="text"/> .	heat of formation	enthalpy of formation				<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	<input type="text"/> metals react with sulphur to form simple sulphides and polysulphides.	Alkaline	Alkaline				<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	Increase in principal quantum number(n) means <input type="text"/> of the atomic radii.	increase	increase				<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	<input type="text"/> is one of the isotopes of hydrogen whose neutron number is zero.	Ordinary hydrogen	Protium				<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	The f-blocks elements are otherwise known as <input type="text"/> .	Inner transition elements	Inner transition elements				<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	The percentage of deuterium in natural occurring hydrogen is <input type="text"/> .	0.000156	0.000156				<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	Inner transition elements are otherwise known as <input type="text"/> .	F-block elements	F-block elements				<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	Hydrogen can be obtained economically as a by-product in the electrolysis of a compound called <input type="text"/> .	brine	Aqueous sodium chloride				<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	The elements which the extra electron enters(n-2)f orbitals are called <input type="text"/> .	F-block elements	Inner transition elements				<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	A consequence of <input type="text"/> is the formation of a chelate ring.	intramolecular hydrogen bonding	intramolecular hydrogen bonding				<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	Halogens have <input type="text"/> electron affinity.	exothermic	exothermic				<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	Hydrogen is discharged at <input type="text"/> electrode during the electrolysis of brine.	Cathode	Cathode				<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	Beryllium, an alkaline earth metal belongs to the <input type="text"/> block elements.	S	S				<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	Fuel cell have advantage over power plants in that its efficiency is <input type="text"/> .	0.75	Seventy five percent				<input type="button" value="eExam"/>

<input type="checkbox"/>	FBQ	The electronic configuration of <input type="text"/> is $1s^2 2s^2 2p^6 3s^1$.	Sodium	Na				eExam
<input type="checkbox"/>	FBQ	The electrons in the same sub shell of a given atom in the ground state occupy different orbitals and will have parallel spin. This is a statement of <input type="text"/> .	Hund's rule	Hund's rule				eExam
<input type="checkbox"/>	FBQ	A hydrogen oxide otherwise known as <input type="text"/> is used as a moderator in nuclear reactors.	Deuterium oxide	D_2O				eExam
<input type="checkbox"/>	FBQ	In nuclear reactors liquid sodium is mainly used as <input type="text"/> .	Coolant	Coolant				eExam
<input type="checkbox"/>	FBQ	<input type="text"/> was the person who arranged elements in order of increasing atomic weight on a line which spiraled around a cylinder from bottom to top.	De Chanourtiot	De Chanourtiot				eExam
<input type="checkbox"/>	FBQ	<input type="text"/> states that as far as possible in a given atom in the ground state, electrons in the same sub shell will occupy different orbitals and will have parallel spins.	Hund's rule	Hund's rule				eExam
<input type="checkbox"/>	FBQ	Which element has the electronic configuration: $[Ar] 3d^8 4s^2$ <input type="text"/>	Nickel	Nickel				eExam
<input type="checkbox"/>	FBQ	The number of nearest neighbours of a metal atom in a body centred cubic is <input type="text"/> .	Eight	8				eExam
<input type="checkbox"/>	FBQ	The atomic number of Rubidium is <input type="text"/> .	thirty seven	37				eExam
<input type="checkbox"/>	FBQ	The periodic table is divided into <input type="text"/> different blocks.	4	Four				eExam
<input type="checkbox"/>	FBQ	The ion-solvent interaction is strong enough to hold a certain fixed number of solvent molecules known as <input type="text"/> .	coordination number	coordination number				eExam
<input type="checkbox"/>	FBQ	Alkali metals react with sulphur to form two types of <input type="text"/> .	sulphides	sulphides				eExam
<input type="checkbox"/>	FBQ	The ion in solution which is surrounded by a number of solvated molecules is called <input type="text"/> .	solvated ion	solvated ion				eExam
<input type="checkbox"/>	FBQ	The name given to the number of nearest neighbours of a metal atom in a lattice is <input type="text"/> .	Coordination number	Coordination number				eExam
<input type="checkbox"/>	FBQ	The polarizing power of a cation is directly proportional to its <input type="text"/> charge.	positive	positive				eExam
<input type="checkbox"/>	FBQ	In the periodic table, period 6 contains <input type="text"/> elements	32	thirty two				eExam
<input type="checkbox"/>	FBQ	The distance from the centre of the nucleus to the point where the electron density is virtually zero is known as <input type="text"/> .	Radius of an atom	Atomic radius				eExam

<input type="checkbox"/>	FBQ	Which group of elements appears in the modern periodic table but not in the Mendeleev's original table? _____.	18	Eighteen					eExam
<input type="checkbox"/>	FBQ	The amount of positive charge felt by the outer electron in an atom is known as _____.	Effective nuclear charge	Effective nuclear charge					eExam
<input type="checkbox"/>	FBQ	The major source for the manufacture of hydrogen is _____.	water	water					eExam
<input type="checkbox"/>	FBQ	The phenomenon in which two forms of hydrogen molecules were formed as a result of differences in the direction of nuclear spin is called _____.	Spin Isomerism	Spin Isomerism					eExam
<input type="checkbox"/>	FBQ	The charge on cations _____ down a group.	remains constant	is constant					eExam
<input type="checkbox"/>	FBQ	_____ is the process of ejecting electrons most readily on exposure of cesium to light.	Photoelectric effect	Photoelectric effect					eExam
<input type="checkbox"/>	FBQ	A mixture of carbon(II)oxide and water is known as _____.	synthesis gas	Water gas					eExam
<input type="checkbox"/>	FBQ	In the modern form of Mendeleev periodic table, elements are arranged in _____ horizontal rows.	7	Seven					eExam
<input type="checkbox"/>	FBQ	The concept of electronegativity was first developed by _____.	Pauling	Pauling					eExam
<input type="checkbox"/>	FBQ	The phenomenon in which a metal is surrounded by water molecules is called _____.	Hydration	Hydration					eExam
<input type="checkbox"/>	FBQ	_____ observed that elements were grouped as triads	J W Dobereiner	Dobereiner					eExam
<input type="checkbox"/>	MCQ	Hydrides of alkali metals react with water to liberate which of these? _____	light	oxygen	hydrogen	ammonia	C		eExam
<input type="checkbox"/>	MCQ	Which of these alkalis has the most stable fluoride? _____	lithium	sodium	potassium	rubidium	A		eExam
<input type="checkbox"/>	MCQ	Why was Newlands law of octaves rejected? _____	Elements were arranged in increasing order of atomic mass	The cycle of repetition shown by elements is similar to octave in music	It could not hold good for elements heavier than calcium	Many radioactive elements had not been discovered then	C		eExam
<input type="checkbox"/>	MCQ	Which of these statements is not true of metallic hydrides? _____	they are volatile	they are deficient in hydrogen	most have metallic appearance	They are conductors of heat	A		eExam
<input type="checkbox"/>	MCQ	Which property was used by Mendeleev to classify elements? _____	atomic number	atomic weight	mass number	chemical properties	D		eExam
<input type="checkbox"/>	MCQ	Which of these alkali metals is the most electropositive? _____	litium	potassium	sodium	caesuim	D		eExam
<input type="checkbox"/>	MCQ	The presence of hydrogen bond in most molecules is responsible for the following EXCEPT _____.	High melting point	Low boiling point	High solubility in water	Very higy boiling point	B		eExam

<input type="checkbox"/>	MCQ	The tendency of an atom to attract towards itself shared electron pair of a bond is known as _____.	electron affinity	electronegativity	atomic attraction	bonding	B	eExam
<input type="checkbox"/>	MCQ	Which of the isotopes of hydrogen is radioactive? _____	Protium	Deutrium	Tritium	Neutrium	C	eExam
<input type="checkbox"/>	MCQ	All these can be used as raw materials for the production of oxygen except _____.	coke	carbon(II)oxide	steam	nitrogen	D	eExam
<input type="checkbox"/>	MCQ	Which of these is not a covalent hydride? _____	Ammonia	Ethyne	Hydrochloric acid gas	NaH	D	eExam
<input type="checkbox"/>	MCQ	The most abundant metal in sea water is called _____.	zinc	sodium	lithium	rubidium	B	eExam
<input type="checkbox"/>	MCQ	Which of these statements is not true of metallic elements? _____	they are ductile	they have lustre	They form acidic oxides	they are malleable	C	eExam
<input type="checkbox"/>	MCQ	Which of the following is not true as we move down a group of s and p-block elements in the periodic table? _____	Effective nuclear charge is relatively constant	Atomic radius increases steadily	The value of principal quantum number increases	Electronegativity also increases	D	eExam
<input type="checkbox"/>	MCQ	Stability of alkali salts depends on which of the following? _____	atomic size of the metal	crystal lattice of the salt formed	intensity of sun rays incident on it	enthalpy of formation of the salt	D	eExam
<input type="checkbox"/>	MCQ	How many electrons are in potassium atom? _____	11	19	23	31	B	eExam
<input type="checkbox"/>	MCQ	How many elements are there in period 6 of the periodic table? _____	32	16	24	8	A	eExam
<input type="checkbox"/>	MCQ	Which of these elements behave partly as the alkaline metals as well as a halogen element? _____	Helium	Fluorine	Hydrogen	Silicon	C	eExam
<input type="checkbox"/>	MCQ	Which of these periods of the peroidic table has only eight elements? _____	1	3	5	7	B	eExam
<input type="checkbox"/>	MCQ	Which of these elements has the electronic configuration of $1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10} 4s^1$? _____	Cu	Zn	Kr	Ca	A	eExam
<input type="checkbox"/>	MCQ	_____ states that as far as possible in a given atom in the ground state, electrons in the same sub shell will occupy different orbitals and will have parallel spins.	Paul Exclusion principles	Hund's rule	Guy's law	Aufbau's principle	B	eExam
<input type="checkbox"/>	MCQ	The alkali metals belong to the _____ elements	d-block	f-block	p-block	s-block	D	eExam
<input type="checkbox"/>	MCQ	Which is of these peroids of the peroidic table has only eight elements? _____	1	3	5	7	B	eExam
<input type="checkbox"/>	MCQ	Increase in principal quantum number(n) means _____ of the atomic radii	shrinking	no change	increase	decrease	C	eExam
<input type="checkbox"/>	MCQ	In the equation for ionization, effective nuclear charge is represented as _____	I	Z*	e	v	B	eExam
<input type="checkbox"/>	MCQ	The ability of an element to participate in a chemical reaction is measured in form of its _____	electron density	electron cloud	ionization energy	nuclear charge	C	eExam
<input type="checkbox"/>	MCQ	_____ is the energy released or absorbed when an electron is added to the gaseous atom in its ground state	electron charge	electronegativity	ionization energy	electron affinity	D	eExam

<input type="checkbox"/>	MCQ	Electron affinity can be affected by all but one of the following _____	effective nuclear charge	atomic radius	mass number	electronic configuration	C	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	Halogens have _____ electron affinity	exothermic	endothermic	positive	zero	A	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	An orbital can have at the most two electrons of opposite spin can be called the _____	aufbau principle	Hendry's law	Hund's rule	Exclusion principle	D	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	In the modern form of Mendeleev periodic table, elements are arranged in _____ horizontal rows.	6	7	8	9	B	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	The positions of K and Ar, Co and Ni do not remain anomalous any longer since _____ is used in arranging the elements.	mass number	atomic weight	atomic number	molecular mass	C	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	The properties of elements are periodic functions of their atomic numbers is the _____	Lecoq de Boisbaudran' law	Modern periodic law	Newlands' law	Mendeleev law	B	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	What property was used by Mendeleev to classify the elements? _____	atomic number	atomic weight	mass number	chemical properties	D	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	_____ used the Law of the Octave to arrange atoms	John Newlands	Meyer	De Chancourtois	Dobereiner	A	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	_____ arranged elements in order of increasing atomic weight on a line which spiraled around a cylinder from bottom to top	John Newlands	Meyer	De Chanourtios	Mendeleev	C	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	_____ and _____ tried to classify elements into periods	De Chancourtois and Dobereiner	Gibbs and Dobereiner	De Chancourtois and Le Chatlier	Graham and Dalton	A	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	_____ is the tendency of an atom to attract towards itself the shared electron pair of a bond in which it is involved	electron affinity	electronegativity	atomic attraction	bonding	B	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	The second ionization energy of alkaline earth metals is less than that of corresponding alkali metals because _____	the alkaline earth metals have higher nuclear charge and are smaller in size	the alkali metals have no nuclear charge	of stability of a closed shell configuration	the alkali metals have higher nuclear charge and are smaller in size	C	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	_____ is the enthalpy change when one mole of crystal lattice is formed from the isolated gaseous ions	hydration energy	lattice energy	ionisation energy	gaseous energy	B	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	Arrange the following metals in an increasing order of their boiling point _____ Li, Cs, Rb, K	Li, Cs, Rb, K	Cs, K, Rb, Li	Rb, Cs, K, Li	Li, K, Rb, C	B	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	Arrange the following alkali metals in order of their atomic number, Na, K, Li,	Na, Li, K	K, Li, Na	Li, K, Na	Li, Na, K	D	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	According to _____, electronegativity is equated to the force of attraction between an atom and the electron separated by a distance equal to the covalent radius of the atom.	Boisbaudran	Alfred-Rochow	Mulliken-Jaffe	Pauling	B	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	Which of these metals can their compounds be used in photography? _____	Ca	Na	K	Cs	C	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	All these are types of metal lattices except _____	hexagonal	cubic close packed	body-centred cubic	hexagonal	D	<input type="checkbox"/> eExam

<input type="checkbox"/>	MCQ	Which of these statements is not true of metallic elements?_____	They form acidic oxides	they are malleable	they have lustre	they are ductile	A	eExam
<input type="checkbox"/>	MCQ	_____is a requirement in the electrolysis of beryllium chloride	Sodium Chloride	calcium	chlorine	beryllium	A	eExam
<input type="checkbox"/>	MCQ	In the extraction of calcium from fused calcium chloride,-----is used as the anode	zinc	graphite	iron	copper	B	eExam
<input type="checkbox"/>	MCQ	All these are uses of Beryllium except _____	used in making atomic fuel containers	used for making hand tools	used for making aircraft	used as window material in X-ray apparatus	C	eExam
<input type="checkbox"/>	MCQ	Calcium oxide is a constituent of all but one of the following_____	mortar	glass	portland cement	barium	D	eExam
<input type="checkbox"/>	MCQ	On exposure to air alkaline earth metals lose their silvery luster because_____	they become dirty	they are over shadowed	a layer of oxide is formed on their surface	they react with carbon (IV) in the atmosphere	C	eExam
<input type="checkbox"/>	MCQ	All the alkaline earth metals form ionic compounds except _____	beryllium	calcium	magnesium	strontium	A	eExam
<input type="checkbox"/>	MCQ	Calcium compounds give out a characteristic ---- flame colouration	brick red	crimson red	apple green	lemon green	A	eExam
<input type="checkbox"/>	MCQ	The hydrogen energies of alkaline earth metal ions are much greater than those of alkali metals because-----	they are larger	they are smaller with decreased cationic charge	they are smaller with increased cationic charge	they have decreased ionic charge	C	eExam
<input type="checkbox"/>	MCQ	Which of these statements is not true of alkaline earth metals?_____	Density decreases down the group from Be to Ca	They are less dense than alkali metals	Density increases after Ra	They have more mass packed into a smaller volume	B	eExam
<input type="checkbox"/>	MCQ	The first ionization energy of alkaline earth metals is more than that of corresponding alkali metals because_____	the alkaline earth metals have higher nuclear charge and are smaller in size	the alkali metals have higher nuclear charge and are smaller in size	the alkali metals have no nuclear charge	the alkaline earth metals have a large size	A	eExam
<input type="checkbox"/>	MCQ	_____is extremely rare and it is a radioactive element	magnesium	radium	beryllium	strontium	B	eExam
<input type="checkbox"/>	MCQ	_____ is the second most abundant metallic element in sea water	copper	magnesium	chloride	lithium	B	eExam
<input type="checkbox"/>	MCQ	Which of these statements is not true of lithium? _____	The salts are less thermally stable than those of the other alkali metals	Lithium reacts spontaneously with water	Lithium does not form solid bicarbonate trioxide or peroxide	Lithium forms stable salts with anions of high charge density	B	eExam
<input type="checkbox"/>	MCQ	The strong cohesive forces present in the lithium atom gives rise to all but one of the following _____	homonuclear bond energy	softness	relatively higher melting point	relatively higher boiling point	B	eExam

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