CHM191

The Bunsen burner regulator can be adjusted to produce a ---------- flame.

\*Non-luminous\*

Electric oven can also be referred to as -------- oven.

\*Drying\*

If corrosive fumes may evolve during evaporation, the process must be carried out in a ---------------.

\*Fume hood\*

The most conventional size for volumetric work is the 250ml -------------------

\*Conical flask\*

Liquid reagent should be taken with the help of -----------------.

\*Droppers\*

The --------------- is the basic unit for the quantity of a substance.

\*Mole\*

The three different techniques in titrimetric analysis are direct, indirect and ---------- titration.

\*Back\*

Aluminium is --------------- reactive than calcium.

\*Less\*

A substance that lose electron is said to be ----------------

\*Oxidized\*

The more ----------------- the potential, the easier the oxidation will be.

\*Positive\*

Heat of -------------- is the amount of heat released or absorbed for a given amount of reactant or products.

\*Reaction\*

The -------------- is used to measure the heat of reaction.

\*Calorimeter\*

-------------- are volumetric analysis involving iodine.

\*Iodimetry\*

KMnO4 is characteristically -------------- in both acid and base media.

\*Coloured\*

In --------------- reactions , oxidation and reduction occurs simultaneously

\*Redox\*

The two balanced ---------------- equations should add up to give redox equations.

\*Half\*

Phenolphthalein can be used in the titration of HCl and NaOH since the titration is between -------------- acid and strong base.

strong

When standard solution directly reacts with the substance being determined, the titration is said to be ------------------- titration.

\*Direct\*

Each types of titrimetric analysis is based on certain types of ------------.

\*Reaction\*

Titration results are best recorded in ----------- decimal places.

\*Two\*

--------------------- of a substance is achieved by titrating the approximate concentration with standard solution of known concentration.

\*Standardization\*

Is it practically possible to prepare a solution of high concentration from a solution of lower concentration?

\*NO\*

A solution contains 2.65g of anhydrous sodium carbonate in 125ml of solution, what is its concentration in mole per dm cube?

\*0.2\*

A standard solution is prepared by weighing a ------- solute and dissolving it in a suitable solvent.

\*Pure\*

The equivalent weight of solute per volume of solution in dm3 is termed ----------

\*Normality\*

Sometimes precipitates are heated to a high temperature to convert it into a ----------- of constant composition.

\*Compound\*

Ash less filter paper leaves little residue on -----------.

\*Ignition\*

For laboratory purpose, filter papers of ---------------- grades are generally made.

\*Three\*

Proper ----------------- of filter paper can increase the rate of filteration.

\*Folding\*

An antidote should be given only to a person that swallow ---------- poisons.

\*Non-corrosive\*

When you are very careful and follow all the laboratory rules of safety ,all accidents in the laboratory will be avoided. True or False

\*False\*

A ------- notebook should be used for laboratory record.

\*Bound\*

Ordinary filtration can be speeded up by the use of ------------------------ filtration.

\*Suction\*

Care must be taken so that the liquid does not --------------- violently while evaporating in a boiling tube.

\*Bump\*

Conical flask is the container in which the ----------- meet and the products is/are formed.

\*Reactants\*

One of the following causes mental confusion

CHCl3

Which of the following is most dangerous to the skin?

HF

Bunsen burner can be adjusted to a moderately high temperature of about \_\_\_\_\_\_\_\_\_.

600 oC

Which of these operations is performed in wet chemical analysis?

Precipitation

The number of moles of solute in one kilogram of solution is called \_\_\_\_\_\_\_\_\_\_

Molality

All pure substance have the following characteristics except \_\_\_\_\_\_\_\_\_\_\_.

Readily soluble in water

There are \_\_\_\_\_\_ types of titrimetric analysis.

Four

Which indicator is suitable for the titration of ethanoic acid and ammonia.

None

Which of the following metals has the least oxidation potential?

Ag

A reaction involving iron and copper ions, which of the ions will be oxidized?

Fe

The reaction is \_\_\_\_\_\_ when the reaction vessel feels warmer.

exothermic

The amount of heat required to raise the temperature of one mole of substances through one degree Celsius is called \_\_\_\_\_\_\_.

Molar heat capacity

Displacement reactions are usually \_\_\_\_\_\_.

oxidized

The most commonly used external indicator for iodine titration is \_\_\_\_\_\_\_\_\_\_\_\_.

starch

It is very difficult to read the meniscus of potassium permanganate solution in burette because \_\_\_\_\_\_\_\_\_\_\_.

Of its intense colouration

Which of these acid is suitable to catalyze permanganate reaction?

Sulphuric acid

Oxidation is defined as the \_\_\_\_\_\_\_\_\_\_\_

Loss of electron

Given the equation, C + H2SO4 → CO2 + SO2 + H2O, the oxidation number of sulphur decrease by \_\_\_\_.

2

The oxidation number of sulphur in H2SO42- is \_\_\_\_\_\_\_\_\_\_\_.

4

H2O2 is oxidized by MnO2 to give \_\_\_\_\_\_\_.

O2

Volumetric method of analysis measures \_\_\_\_\_\_.

Volume of a gas or volume of solution of unknown concentration

The concept of \_\_\_\_\_\_\_\_ method is used in the preparation of a solution of lower concentration from higher concentration

Dilution

Which of the following is true of molar and molal concentration of sodium carbonate?

Molal concentration is higher

Precipitation is common to \_\_\_\_\_\_\_\_ analysis.

Gravimetric

What should be given to a person that swallowed corrosive poison as soon as possible?

Calcium hydroxide solution

Burns caused by bromine can be treated with \_\_\_\_\_\_\_

Ammonia

Burns are mostly caused by \_\_\_\_\_\_.

Hot equipment and reagents

What is the molar concentration of a solution containing 2.5g of potassium hydroxide in 200cm3 of solution. (K = 39, H=1,O=16)

0.22mol/dm3

In a titration, 25cm3of an impure anhydrous sodium trioxocarbonate (iv) containing 5.0g in 1dm3 of solution was neutralized by 22.20cm3of 0.10mol/dm3 of HCl, find the mass concentration of the trioxocarbonate (iv).

4.71g/dm3

A molar solution of caustic soda is prepared by dissolving \_\_\_\_\_\_\_\_\_\_\_\_.

40g NaOH in 1000g of water

What volume of 0.5M sulphuric acid will exactly neutralize 20cm3 of 0.1M NaOH solution?

2.0 cm3

In which type of titration is two standard solution needed?

Back

The volume of acid used is usually \_\_\_\_\_\_

Average titre value

All electron donors are normally \_\_\_\_\_.

Reducing agent

At what reaction condition will the absolute value of enthalpy equals the absolute value of heat of reaction?

Constant pressure