

eExam Question Bank

Coursecode:

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<input type="checkbox"/>	Question Type	Question	A	B	C	D	Answer	Remark
<input type="checkbox"/>	FBQ	Sodium metal [] on contact with water.	explode	explodes				<input type="checkbox"/> eExam
<input type="checkbox"/>	FBQ	Aqueous ammonia should be used only in the [] hood.	fume					<input type="checkbox"/> eExam
<input type="checkbox"/>	FBQ	NaOH is [] and can eat away living tissues.	caustic					<input type="checkbox"/> eExam
<input type="checkbox"/>	FBQ	[] are used to heat small quantities of substances to very high temperatures.	crucibles					<input type="checkbox"/> eExam
<input type="checkbox"/>	FBQ	A medicine dropper is calibrated by counting the number of drops it produces to make up a [] .	millilitre					<input type="checkbox"/> eExam
<input type="checkbox"/>	FBQ	[] flasks are used to prepare solution of a specific concentration.	volumetric					<input type="checkbox"/> eExam
<input type="checkbox"/>	FBQ	All edges of glass tubing must be fine-polished to round off the [] edges.	sharp					<input type="checkbox"/> eExam
<input type="checkbox"/>	FBQ	Never place the litmus paper [] into the solution.	directly					<input type="checkbox"/> eExam
<input type="checkbox"/>	FBQ	[] is the container used to keep sample dry.	A desiccator	Desiccator				<input type="checkbox"/> eExam

<input type="checkbox"/>									
<input type="checkbox"/>	FBQ	A simple oil bath can be made by heating a <input type="text"/> of oil on a hot plate.	dish						eExam
<input type="checkbox"/>	FBQ	The size of steam bath can be controlled by adding or removing <input type="text"/> .	rings						eExam
<input type="checkbox"/>	FBQ	Heating mantles must be used with variable <input type="text"/> .	transformers						eExam
<input type="checkbox"/>	FBQ	The heating mantle is safe because it does not produce <input type="text"/> .	flames	flame					eExam
<input type="checkbox"/>	FBQ	Below the <input type="text"/> point, solvent will be seen running back into the flask.	condensation						eExam
<input type="checkbox"/>	FBQ	<input type="text"/> is the process of boiling reactants while cooling the vapour and returning it to the flask as liquid.	reflux						eExam
<input type="checkbox"/>	FBQ	For micro-scale work, a Pasteur dropping <input type="text"/> can be used instead of glass funnel.	pipette						eExam
<input type="checkbox"/>	FBQ	Filtration involves the separation of <input type="text"/> solid materials from a liquid.	insoluble	suspended					eExam
<input type="checkbox"/>	FBQ	When heating beakers with a burner, <input type="text"/> should be placed between the vessel and the flame.	wire gauze						eExam
<input type="checkbox"/>	FBQ	For temperature above 100°C, <input type="text"/> baths are generally used.	oil						eExam
<input type="checkbox"/>	FBQ	A stirrer hot plate keeps the <input type="text"/> at a constant temperature while stirring.	solution						eExam

<input type="checkbox"/>									
<input type="checkbox"/>	FBQ	Powerful pumps provide good heat <input type="text"/> and optimum temperature accuracy.	exchange						eExam
<input type="checkbox"/>	FBQ	Bunsen burner is used for heating, sterilisation and <input type="text"/> -	combustion						eExam
<input type="checkbox"/>	FBQ	Most laboratory <input type="text"/> are equipped with multiple gas nozzles connected to a central gas source.	benches						eExam
<input type="checkbox"/>	FBQ	The most common methods of lighting the burner are using a match on a <input type="text"/> lighter.	spark/flame	flame					eExam
<input type="checkbox"/>	FBQ	A Bunsen burner was named after <input type="text"/> Bunsen.	Robert						eExam
<input type="checkbox"/>	FBQ	Proper selection of a suitable type of stirrer requires that you certain <input type="text"/> variables.	application						eExam
<input type="checkbox"/>	FBQ	The two types of magnetic stirrers are mechanical and <input type="text"/> -	electronic						eExam
<input type="checkbox"/>	FBQ	The most important applications of magnetic stirring is stirring in a <input type="text"/> system.	closed						eExam
<input type="checkbox"/>	FBQ	The viscosity of most liquids varies <input type="text"/> with temperature.	inversely						eExam
<input type="checkbox"/>	FBQ	Modern digital weighing scales have many types, styles and <input type="text"/> ranges.	weigh						eExam
<input type="checkbox"/>	FBQ	Modern digital weighing scales <input type="text"/> results very precisely.	display						eExam

<input type="checkbox"/>									
<input type="checkbox"/>	FBQ	Precision and <input type="text"/> were still missing with the use of spring balance.	accuracy						eExam
<input type="checkbox"/>	FBQ	The conventional <input type="text"/> scale was the oldest weighing machine invented by man.	wooden	iron					eExam
<input type="checkbox"/>	FBQ	Modern weighing machines constitute of electronic circuits, digital <input type="text"/> _, electronic and motion sensors.	chips						eExam
<input type="checkbox"/>	FBQ	Hydraulic scales make use of liquid <input type="text"/> technique to weigh the mass of heavy objects.	displacement						eExam
<input type="checkbox"/>	FBQ	Hydraulic scales are used mainly in the laboratories and <input type="text"/> units.	industrial						eExam
<input type="checkbox"/>	FBQ	Use <input type="text"/> balance if the experiment calls for accurate measurement of mass.	analytical						eExam
<input type="checkbox"/>	FBQ	The <input type="text"/> balance is more accurate than the top-loading balance.	analytical						eExam
<input type="checkbox"/>	FBQ	The top-loading balance <input type="text"/> displays a mass reading in grams.	digitally						eExam
<input type="checkbox"/>	FBQ	The uncertainty in a single reading on the top-loading balance is <input type="text"/> .	0.05g						eExam
<input type="checkbox"/>	FBQ	To avoid <input type="text"/> of moisture, refrigerated materials must be allowed to come to room temperature before the container is opened.	condensation						eExam

<input type="checkbox"/>									
<input type="checkbox"/>	FBQ	Electronic analytical balance, has an <input type="text"/> calibration system based on the applied load.	internal						eExam
<input type="checkbox"/>	FBQ	The power of the balance could be turned on and the balance allowed to <input type="text"/> for at least one hour before proceeding with the calibration.	equilibrate						eExam
<input type="checkbox"/>	FBQ	Hysteresis of the <input type="text"/> parts occurs during weighing.	mechanical						eExam
<input type="checkbox"/>	FBQ	Balance drift can be present without the <input type="text"/> being aware of it.	operators	operator					eExam
<input type="checkbox"/>	FBQ	Drift is one of the most common <input type="text"/> —.	errors	error					eExam
<input type="checkbox"/>	FBQ	When a balance is moved, it must be allowed to adjust to the temperature of its new environment and be <input type="text"/> —.	re-calibrated						eExam
<input type="checkbox"/>	FBQ	To make sure that the <input type="text"/> of the air in the balance does not change, keep the balance doors closed.	temperature						eExam
<input type="checkbox"/>	FBQ	The weighing balance is levelled if the bubble in the level <input type="text"/> is in the centre.	indicator						eExam
<input type="checkbox"/>	FBQ	Do not weigh cold objects on the balance because cold object usually give <input type="text"/> readings.	high	large					eExam
<input type="checkbox"/>	FBQ	Weighing hot objects gives low reading due to <input type="text"/> of hot air.	buoyancy						eExam
<input type="checkbox"/>	FBQ	Ensure that you do not <input type="text"/> the balance.	overload						eExam

<input type="checkbox"/>									
<input type="checkbox"/>	FBQ	After use, wash the Erlenmeyer flasks and do final rinsing with <input type="text"/> water.	deionised						eExam
<input type="checkbox"/>	FBQ	The following items must be taken to the weighing room: Erlenmeyer flask, <input type="text"/> and pen.	report sheet						eExam
<input type="checkbox"/>	FBQ	Turn the analytical balance on and allow it to <input type="text"/> —.	calibrate						eExam
<input type="checkbox"/>	FBQ	Read the balance and record the <input type="text"/> directly onto your record sheet.	mass						eExam
<input type="checkbox"/>	FBQ	To use the Top-loading balance, check the balance for dust, <input type="text"/> off any that you see.	brushing						eExam
<input type="checkbox"/>	FBQ	Throughout the weighing process, you should protect the Erlenmeyer flasks from your hand with a <input type="text"/> towel.	paper						eExam
<input type="checkbox"/>	FBQ	A simple distillation involves a <input type="text"/> equilibration between the liquid and vapour.	single	mono					eExam
<input type="checkbox"/>	FBQ	Azeotropes <input type="text"/> the boiling behaviour of pure liquids.	mimic	copy					eExam
<input type="checkbox"/>	MCQ	In the laboratory, the most common measuring instrument is _____.	Graduated cylinder	Volumetric flask	Burette	Volumetric beaker	A		eExam
<input type="checkbox"/>	MCQ	When alcohol is dehydrated, product x is formed. To prevent x from reforming back to alcohol, x is preferably removed by _____.	crystallization	distillation	filtration	decantation	B		eExam
<input type="checkbox"/>	MCQ	Elimination of a molecule of water in alcohol results in Pi-bond formation of _____	Alkynes and alkenes	Alkynes, alkenes and alkanes	Alkynes and alkanes	Alkenes and alkanes	A		eExam

<input type="checkbox"/>	MCQ	Elimination reaction of an alcohol is called _____.	addition	substitution	catalysis	dehydration	D	eExam
<input type="checkbox"/>	MCQ	Calcium chloride is mostly used as _____.	Drying agent	Dehydrating agent	Reducing agent	Emulsifying agent	A	eExam
<input type="checkbox"/>	MCQ	In the reaction of KMnO_4 and ethene, the KMnO_4 is _____.	reduced	oxidized	neutralized	hydrolysed	A	eExam
<input type="checkbox"/>	MCQ	Bromine water is reacted with I. Butane II. Butene III. Butyne. In which of the/these reactions will bromine water be decolourised?	I and III	II,III	I,II	I,II,III	B	eExam
<input type="checkbox"/>	MCQ	Which of the following reactions (1) substitution (2) addition (3) reaction with sulphuric acid, is/are not common to alkanes?	2&3	1&3	1&2	1,2&3	A	eExam
<input type="checkbox"/>	MCQ	Compounds with at least one carbon-to-carbon triple bond are called	alkenes	alkanes	alkynes	aromatic	C	eExam
<input type="checkbox"/>	MCQ	Cyclic hydrocarbons are aliphatic compound that have carbon atoms linked _____.	By branch	In a closed polygon	Without ring	In a circle	B	eExam
<input type="checkbox"/>	MCQ	Hydrocarbons consisting of carbon linked in a single chain are chain aliphatic, those linked by branched chain are called _____.	Chain aliphatic	Cyclic hydrocarbon	Branched aliphatic	Cyclic aliphatic	A	eExam
<input type="checkbox"/>	MCQ	The simplest formula for representing hydrocarbon is called _____ formula.	empirical	structural	molecular	homologous	A	eExam
<input type="checkbox"/>	MCQ	$\text{BrCH}_2\text{CH}_2\text{Br}$ represent _____.	1,4-dibromoethane	1,2-dibromoethane	1,1-dibromoethane	2,3-dibromoethane	B	eExam
<input type="checkbox"/>	MCQ	One of the reasons for the vast array of organic compounds is the phenomenon of _____.	Allotropy	Azeotropes	Carbonation	Isomerism	D	eExam
<input type="checkbox"/>	MCQ	Isomers have different physical properties and _____ chemical properties.	different	The same	Almost the same	All of the above	A	eExam
<input type="checkbox"/>	MCQ	To write the _____ formula of hydrocarbon, the hydrogen atoms are grouped with their carbon atom.	empirical	structural	skeletal	condensed	D	eExam

<input type="checkbox"/>	MCQ	Alkanes are saturated hydrocarbons, cycloalkanes are _____ hydrocarbons.	saturated	unsaturated	Straight chain	None of the above	A	eExam
<input type="checkbox"/>	MCQ	Some solids begin to _____ a few degrees below their melting point.	decompose	sweat	sublime	expand	B	eExam
<input type="checkbox"/>	MCQ	The shrinkage of a solid being heated is called _____.	sintering	recoiling	eutectic	contraction	A	eExam
<input type="checkbox"/>	MCQ	When liquid is cooled less than its freezing point, it _____.	does solidify	expands	Does not solidify	Does not melt	C	eExam
<input type="checkbox"/>	MCQ	Supercooling occurs when a liquid is cooled _____ its freezing point.	to	below	above	within	B	eExam
<input type="checkbox"/>	MCQ	The _____ of a solid is the temperature at which the liquid and solid phases are in equilibrium.	Melting point	Freezing point	Melting point and freezing point	None of the above	C	eExam
<input type="checkbox"/>	MCQ	Which of these; ethanol and methanol, has higher boiling point and why?	Ethanol, it is polar	Methanol, has lower molecular mass	Methanol, it is non-polar	Ethanol, it has higher molecular mass	D	eExam
<input type="checkbox"/>	MCQ	What effect do impurities have on the melting point of solids?	Raises it	Reduces it	No effect	It depends on the solid	B	eExam
<input type="checkbox"/>	MCQ	A decrease in atmospheric pressure, _____ the boiling point of a substance	increases	has no effect	reduces	None of the above	C	eExam
<input type="checkbox"/>	MCQ	1. Polarity 2. Branching 3. Molecular weight. Which of the above features reduces the boiling point of a compound?	2 only	3 only	1 only	1 & 2 only	A	eExam
<input type="checkbox"/>	MCQ	1. Polarity 2. Branching 3. Molecular weight. Which of the above features increases the boiling point of a compound?	1&2	1&3	2&3	3 only	B	eExam
<input type="checkbox"/>	MCQ	What effect does formal charge in a molecule have on the boiling point of a substance?	Increases it.	No effect.	Reduces it.	It can either reduce or increase it.	A	eExam
<input type="checkbox"/>	MCQ	Calculate the boiling point of ethanol at 700mmHg when its normal points at 760mmHg is known to be 79oC, assuming the boiling point will change by 0.5oC for each 10mmHg.	75.3oC	76.0oC	75.0oC	77.0oC	B	eExam

<input type="checkbox"/>	MCQ	1. Polarity 2. Branching 3. Molecular weight. Which of the listed factors (1 - 3) affects the boiling points of substances?	1 and 3	2 and 3	1,2 and 3	1 and 2	C	eExam
<input type="checkbox"/>	MCQ	Mercury is traded by the "flask" with mass 34kg. What is the volume of flask of mercury, if the density of mercury is 13kg/ml.	2.62ml	0.382ml	21ml	47ml	A	eExam
<input type="checkbox"/>	MCQ	A bottle contains 5dm ³ of metallic solid, the total mass of the bottle is 35g. The empty bottle weighs 14g. What is the density of the solid?	9.8g/ dm ³	4.2 g/ dm ³	4.59 g/ dm ³	4.0 g/ dm ³	B	eExam
<input type="checkbox"/>	MCQ	A cylinder is filled to the 40ml mark with oil, the masses of the cylinder before and after the addition of the oil are 124g and 159g. Calculate the density of the oil.	0.915g/ml	0.862g/ml	0.951g/ml	0.975g/ml	D	eExam
<input type="checkbox"/>	MCQ	Solvents for extraction should have reasonably high boiling point.	CORRECT	NOT CORRECT	MAYBE	NOT SURE	B	eExam
<input type="checkbox"/>	MCQ	Can form emulsion, easy to dry and remove heat hazard, suspected carcinogen are all properties of _____.	ETHYL ACETATE	CHLOROFORM	METHYLENE DICHLORIDE	ETHYL ETHER	B	eExam
<input type="checkbox"/>	MCQ	_____ is the most common extraction solvent.	methylene chloride	ethyl ether	dichloromethane	isobutyric acid	B	eExam
<input type="checkbox"/>	MCQ	In technical sense, extraction is based on the principle of the equilibrium _____ of a substance.	distribution	redistribution	management	adjustment	A	eExam
<input type="checkbox"/>	MCQ	Which of the following separation method is sufficient to separate an organic substance from a mixture?	Isolation	Extraction	Purification	Dissolution	B	eExam
<input type="checkbox"/>	MCQ	Which of these separation techniques yield impure products?	Extraction	Recrystallization	Distillation	Crystallization	A	eExam
<input type="checkbox"/>	MCQ	The reciprocal of (solubility in hot solvent ÷ solubility in cold solvent) is same as _____.	Recovery of solid	Loss in cold solvent	Loss in hot solvent	Recovery of solvent	B	eExam
<input type="checkbox"/>	MCQ	Which of the following is the solubility in water increasing?	n-hexanol→n-heptanol→n-pentanol	n-pentanol→n-hexanol→n-heptanol	n-heptanol→n-hexanol→n-pentanol	n-hexanol→n-heptanol→n-pentanol	C	eExam

<input type="checkbox"/>								
<input type="checkbox"/>	MCQ	Which of the following is more soluble in water?	n-pentanol	n-butanol	n-heptanol	n-hexanol	B	eExam
<input type="checkbox"/>	MCQ	Which of the following is less soluble in polar solvent?	n-butanol	n-pentanol	n-heptanol	n-hexanol	C	eExam
<input type="checkbox"/>	MCQ	In polar solvent, the solubility of alcohol is _____ proportional to the number of carbons per polar functional group.	inversely	directly	partially	not	A	eExam
<input type="checkbox"/>	MCQ	In selecting a good re-crystallization solvent, one should consider _____.	Expense and flammability	Toxicity, solubility and expense	Flammability, expense and solubility	Expense, toxicity and flammability	D	eExam
<input type="checkbox"/>	MCQ	The most basic technique for purifying organic solid is _____.	crystallization	emulsification	salting	decantation	A	eExam
<input type="checkbox"/>	MCQ	The purity of solid can be verified by checking the _____.	Molar mass	Standard density	Melting points	Crystallization	C	eExam
<input type="checkbox"/>	MCQ	Solids are purified by _____.	crystallisation	dehydration	drying	filtration	A	eExam
<input type="checkbox"/>	MCQ	To separate a mixture of miscible liquids with different boiling points, you employ _____.	Distillation	Fractional distillation	Distillation and Fractional distillation	Recrystallisation	C	eExam
<input type="checkbox"/>	MCQ	In distillation process, the distilled volume is usually measured in _____.	cm ³	dm ³	ml	None of the above	C	eExam
<input type="checkbox"/>	MCQ	Which of the following is true?	Methanol is flammable	Do not distil the flask to dryness	All glass wares must be clamped in space	All of the above	D	eExam
<input type="checkbox"/>	MCQ	Which of the following has the highest boiling point?	Bromoform	Water	Octane	Toluene	A	eExam
<input type="checkbox"/>	MCQ	Simple distillation can be used to separate the mixture of t-butyl alcohol, 2-propanol and ethanol.	That's true	not true	Not sure	None of the above	B	eExam
<input type="checkbox"/>	MCQ	Find the mole fraction of ethanol in the mixture of 95.0g of ethanol, CH ₃ CH ₂ OH and 5.0g water.	0.88	2.07	0.12	0.28	A	eExam
<input type="checkbox"/>	MCQ	Calculate the mole fraction of water in the mixture of 95.0g ethanol, CH ₃ CH ₂ OH and 5.0g water.	0.88	2.07	0.12	0.28	C	eExam

<input type="checkbox"/>								
<input type="checkbox"/>	MCQ	The normal boiling point of a liquid is the temperature at which the vapour pressure of a liquid equal _____ pressure.	applied	atmospheric	partial	boiling	B	eExam
<input type="checkbox"/>	MCQ	The boiling point of a liquid is a measure of its _____	purity	volatility	vapour pressure	boiling temperature	B	eExam
<input type="checkbox"/>	MCQ	Which separation method purify a compound by separating it from less volatile material?	crystallization	filtration	decantation	distillation	D	eExam
<input type="checkbox"/>	MCQ	At any given temperature, a liquid is in equilibrium with its _____.	vapour	density	solid	volume	A	eExam
<input type="checkbox"/>	MCQ	Distillation is based on the difference in _____ of mixture.	density	boiling point	melting point	volatility	B	eExam

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