

eExam Question Bank

Coursecode:

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Search:

<input type="checkbox"/>	Question Type	Question	A	B	C	D	Answer	Remark
<input type="checkbox"/>	FBQ	<input type="text"/> type of error cannot be avoided.	Random error	Indeterminate error				<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	<input type="text"/> help in eliminating the possible sources of contamination and sample degradation that could lead to sample destruction, and ensure the homogeneity of samples.	Treatment of sample	Preparation of sample				<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	Observing sudden colour change in the indicator is one of the methods of detecting completion of a <input type="text"/> .	Titration reaction					<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	The decrease in power of the beam of a monochromatic radiation when it passes through absorbing specie depends upon the <input type="text"/> of the absorber and the length of the path transverse by the beam.	Concentration					<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	The device in which electrolysis of solution takes place is known as <input type="text"/> .	Electrochemical cell					<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	In redox titration <input type="text"/> measures in voltage the concentration of species in solution.	Potentiometer					<input type="button" value="eExam"/>

<input type="checkbox"/>									
<input type="checkbox"/>	FBQ	The part of the component of a spectrometer that selects a narrow band of wavelength from the source spectrum is called <input type="text"/> —.	Monochromator						eExam
<input type="checkbox"/>	FBQ	<input type="text"/> reduce the water content of a sample material	Concentration						eExam
<input type="checkbox"/>	FBQ	In redox titration, half-cell reaction occur at each <input type="text"/> —.	Electrode						eExam
<input type="checkbox"/>	FBQ	<input type="text"/> is a compound with a physical property (colour) which changes abruptly near the equivalence end point.	Indicator						eExam
<input type="checkbox"/>	FBQ	<input type="text"/> type of error are indicated by small differences in successive measurements made by the same analyst under almost identical experimental conditions.	Random error	Indeterminate error					eExam
<input type="checkbox"/>	FBQ	<input type="text"/> is a method of sampling employed when population is heterogeneous with respect to the variables (parameters) under study.	Stratified sampling						eExam
<input type="checkbox"/>	FBQ	A typical electrolytic cell is made up of electrodes, salt bridge and <input type="text"/> —.	Electrolyte						eExam
<input type="checkbox"/>	FBQ	The process by which precipitates revert to the colloidal state is called <input type="text"/> —.	Peptization						eExam
<input type="checkbox"/>	FBQ	<input type="text"/> is defined as the maximum amount of an acid or base that can be added to a buffer system without causing a change or appreciable change in pH of a system.	Buffer capacity						eExam

<input type="checkbox"/>									
<input type="checkbox"/>	FBQ	The process by which a precipitate is allowed to stand in the solution from which it was precipitated is known as <input type="text"/>	Digestion of precipitate	Digestion					eExam
<input type="checkbox"/>	FBQ	pH of 7.0 means the solution is <input type="text"/>	Neutral						eExam
<input type="checkbox"/>	FBQ	In volatilization gravimetric analysis, the substance to be determined is separated in a <input type="text"/> form from the remainder of the sample.	Gas						eExam
<input type="checkbox"/>	FBQ	On the pH scale acidity increases from <input type="text"/> TO <input type="text"/>	0.0, 6.00	0.0, 6.00					eExam
<input type="checkbox"/>	FBQ	Round off 9.43 to two significant figures <input type="text"/>	9.4						eExam
<input type="checkbox"/>	FBQ	<input type="text"/> affect the precision and accuracy of a measured quantity thereby raising questions on the integrity of the reported values.	Errors	Error					eExam
<input type="checkbox"/>	FBQ	<input type="text"/> is a process by which the precise concentration of a solution is determined.	Standardisation						eExam
<input type="checkbox"/>	FBQ	<input type="text"/> involves converting sample material in solid form to solution.	Dissolution						eExam
<input type="checkbox"/>	FBQ	The potential of hydrogen in a system is known as <input type="text"/>	pH						eExam
<input type="checkbox"/>	FBQ	<input type="text"/> can be described as the operations involved in procuring a laboratory size that is a true representative of "a whole lot" for a particular analytical exercise.	Sampling						eExam

<input type="checkbox"/>									
<input type="checkbox"/>	FBQ	Precipitates are treated to have it in a form suitable for weighing by <input type="text"/> —.	Drying	heating					eExam
<input type="checkbox"/>	FBQ	<input type="text"/> types of sample provide more representative sampling of heterogeneous matrices in which the composition of the analyte of interest may vary over a period of time and or space.	Composite samples						eExam
<input type="checkbox"/>	FBQ	<input type="text"/> measures how closely the data cluster about the mean.	Standard deviation						eExam
<input type="checkbox"/>	FBQ	<input type="text"/> is a solution that resists changes in pH when small amount of acid or base is added or when dilution occurs.	Buffer solution						eExam
<input type="checkbox"/>	FBQ	<input type="text"/> involves heating the materials at high temperature $\$ \$ \left(400 - 700^{\circ} \right) \$ \$$ In a muffle furnace until it is turned to ashes.	Dry ashing	Ashing					eExam
<input type="checkbox"/>	FBQ	<input type="text"/> is defined as the degree of agreement between replicate measurements of the same quantity.	Precision						eExam
<input type="checkbox"/>	FBQ	At elevated/high temperature in a fluid system the pH turns out to be <input type="text"/> —.	High						eExam
<input type="checkbox"/>	FBQ	pH of 5.4 means the solution is <input type="text"/> —.	Acidic						eExam
<input type="checkbox"/>	FBQ	Precipitates are washed after filtration so as to remove any <input type="text"/> —.	Coprecipitate impurities	Impurities					eExam

<input type="checkbox"/>									
<input type="checkbox"/>	FBQ	The amount of radiation absorbed by a sample is determined by <input type="text"/> law.	Beer's						eExam
<input type="checkbox"/>	FBQ	The region of the electromagnetic spectrum that extends from about 200 – 380 nm is known as <input type="text"/> —.	Near ultraviolet region						eExam
<input type="checkbox"/>	FBQ	The statistical tool used most frequently to compare the mean value from experimental procedure is known as <input type="text"/> —.	Student's t Test	Student's t Test					eExam
<input type="checkbox"/>	FBQ	Alkalinity increases from <input type="text"/> on the pH scale.	8.0 – 14.0	Eight to fourteen					eExam
<input type="checkbox"/>	FBQ	During digestion of a precipitate, the process whereby precipitates carry down from the solution other constituent that cause contamination of the precipitate is called <input type="text"/> —.	Coprecipitation						eExam
<input type="checkbox"/>	FBQ	The type of titration used in determining the concentration of analyte which precipitate with the anion or titrant is known as <input type="text"/> .	Precipitation titration						eExam
<input type="checkbox"/>	FBQ	<input type="text"/> is a method of sampling in which each item of the population has equal chance of being included in the samples.	Random sampling						eExam
<input type="checkbox"/>	FBQ	The tendency to oxidise or reduce depends on the <input type="text"/> of a substance.	Reduction potential						eExam
<input type="checkbox"/>	FBQ	The actual point when a reaction is observed to be complete is known as <input type="text"/> —.	End point						eExam

<input type="checkbox"/>									
<input type="checkbox"/>	FBQ	<input type="text"/> is an analytical method which is used to determine larger number metals that form soluble salt or slightly dissociated complexes.	Complexometric titration						eExam
<input type="checkbox"/>	FBQ	<input type="text"/> is an analytical method that is based upon the measurement of the weight of known composition.	Gravimetry	Gravimetric analysis					eExam
<input type="checkbox"/>	FBQ	The sample container which holds the sample to be analysed by a spectrometer must be <input type="text"/> in the wavelength region being measured.	Transparent						eExam
<input type="checkbox"/>	FBQ	Ultraviolet region extends from <input type="text"/> —.	10 to 380 nm						eExam
<input type="checkbox"/>	FBQ	The end point of a <input type="text"/> can be determined by the use of electrode to measure potential, or use of self indicator or use of starch indicator.	Redox titration	Oxidation-reduction titration					eExam
<input type="checkbox"/>	FBQ	The difference between the equivalence point and end point is known as <input type="text"/> —.	Titration error						eExam
<input type="checkbox"/>	FBQ	<input type="text"/> is the purest form of reagent which is used to prepare a standard solution.	Primary standard						eExam
<input type="checkbox"/>	FBQ	In titration reaction, the end point should coincide with the <input type="text"/> .	Equivalence point						eExam
<input type="checkbox"/>	FBQ	<input type="text"/> is an analytical technique that deals with reactions between measured volumes of a reagent against the test substance called analyte in a stoichiometric manner	Volumetric analysis						eExam

<input type="checkbox"/>									
<input type="checkbox"/>	FBQ	<input type="text"/> brought to the laboratory require further treatment before analysis commences.	Sample	Samples					eExam
<input type="checkbox"/>	FBQ	Samples composed of mixture of grab samples collected from different points simultaneously or as nearly so as possible is known as <input type="text"/> —.	Integrated samples						eExam
<input type="checkbox"/>	FBQ	The nature of a <input type="text"/> must be the same with that of the population and must remain so throughout the analytical exercise.	Sample						eExam
<input type="checkbox"/>	FBQ	Errors inherent in the analytical method or procedure used is known as <input type="text"/> —.	Methodic error						eExam
<input type="checkbox"/>	FBQ	Mathematical error in the calculation of measurements or prejudice in estimating measurement is known as <input type="text"/> —.	Operative error						eExam
<input type="checkbox"/>	FBQ	<input type="text"/> occurs when faulty equipment, weight as well as glassware used are not well calibrated or wrongly calibrated.	Instrumental error						eExam
<input type="checkbox"/>	FBQ	Errors that can be determined and corrected are known as <input type="text"/> —.	Systematic error						eExam
<input type="checkbox"/>	FBQ	<input type="text"/> is termed controllable and uncontrollable variations observed when comparing the measured values to the true value.	Error						eExam
<input type="checkbox"/>	MCQ	Molecules can absorb radiation by _____.	Molecular transition	Spectroscopic transition	Vibrational transition	Excitation transition	C		eExam

<input type="checkbox"/>	MCQ	The amount of radiation absorbed by a sample can be _____.	Related quantitatively to the concentration of analyte in solution	Transmitted without absorption	None of the above	All of the above	B	eExam
<input type="checkbox"/>	MCQ	When atoms or molecules absorb radiation, _____.	The atoms or molecules moves to lower internal energy level	There will be change in state in the number of the atoms or molecules	The atoms or molecules moves to higher internal energy level	The atoms or molecules remains in the same energy level	C	eExam
<input type="checkbox"/>	MCQ	_____ is an example of a metal-ion indicator.	Methyl orange	Erichrome black T	Starch	Ferroin	B	eExam
<input type="checkbox"/>	MCQ	Upon absorbing a photon of energy from radiation, atoms or molecules _____.	Loss back the photon	Moves to the transit state	Moves to the lowest electronic energy state	Moves to higher electronic energy state	D	eExam
<input type="checkbox"/>	MCQ	EDTA is hexaprotic H_6Y^{2+} because _____.	The number of acidic hydrogen atom lost upon complete metal complex formation is six	The number of carboxyl lost upon complete metal complex formation is six	The number of ammonium lost upon complete metal complex formation is six	The number of nitrogen atom lost upon complete metal complex formation is six	A	eExam
<input type="checkbox"/>	MCQ	_____ is a form of volumetric analysis in which the formation of a coloured complex is used to indicate the end point of the titration.	Precipitation titration	Complexometric titration	Acid-base titration	Oxidation-reduction titration	B	eExam
<input type="checkbox"/>	MCQ	Which of the following options below shows the correct arrangement of a spectrometer?	Read out, detector, sample, monochromator and source.	Sample, source, monochromator, detector and read-out.	Source, monochromator, sample, detector and read-out.	Monochromator, source, sample, detector and read-out.	C	eExam
<input type="checkbox"/>	MCQ	The region of the electromagnetic spectrum that extends from 380 nm to about 780 nm is called _____.	Macrowave region	Visible region	Ultraviolet region	Infrared region	B	eExam
<input type="checkbox"/>	MCQ	In the titration of a strong acid against strong bases, before equivalence point _____.	The pH is determined by the excess of OH in the solution.	The pH is determined by dissociation of water.	The pH is determined by the excess H ⁺ in the solution.	The pH is determined by equilibrium between OH and H ⁺ .	A	eExam
<input type="checkbox"/>	MCQ	Good solvent for many organics is _____.	Aqueous solutions	Water	Weak mineral acids	Strong mineral acids	D	eExam
<input type="checkbox"/>	MCQ	_____ law states that "when a monochromatic radiation passes through absorbing specie, the power of the beam is progressively decreased as more energy is absorbed by the particle".	Rutherford's	Schroedinger's	Beer's	Hess	C	eExam

<input type="checkbox"/>	MCQ	_____ consist of a mixture of a weak acid and its salt, or a weak base and its salt.	Buffer solution	Double salt	Acidic salt or basic salt	None of the above	A	eExam
<input type="checkbox"/>	MCQ	The pH of most solution is in the range of _____.	0.0 – 0.7	0.0 -14.0	0.7 – 14.0	8.0 – 0.0	B	eExam
<input type="checkbox"/>	MCQ	When a precipitate is heated, it is continued until _____.	It is completely dry	It decomposes	It is ashed	A constant weight is achieved	D	eExam
<input type="checkbox"/>	MCQ	Filterability of a precipitate obtained from gravimetric analysis is enhanced if the precipitate _____.	Exist in solid form	Is devoid of impurities	Is soluble	Contain larger crystals	D	eExam
<input type="checkbox"/>	MCQ	What happens when the precipitate obtain from gravimetric analysis is allowed to stand?	The precipitate dissolves	Larger crystals grow at the expense of small crystals	Smaller crystals grow at the expense of larger crystals	The quantity of precipitate obtained increases	B	eExam
<input type="checkbox"/>	MCQ	Which of the following is not a step of gravimetric analysis?	Drying	Precipitation	Extraction	Washing	C	eExam
<input type="checkbox"/>	MCQ	Peptization of a precipitate can occur by _____.	Washing the obtained precipitate with an electrolyte	Heating the precipitate	Filtering the precipitate	Washing the precipitate with water	D	eExam
<input type="checkbox"/>	MCQ	In gravimetry, precipitation occurs through _____.	Supersaturation, crystal growth and nucleation	Crystal growth, nucleation and supersaturation	Nucleation, supersaturation, and crystal growth	Supersaturation, nucleation and crystal growth	D	eExam
<input type="checkbox"/>	MCQ	The commonest way of removing interferences in gravimetry is by _____.	Dissolution	Introducing reagent that selectively mask the interfering substances	Dissolution	Heating the sample to remove interfering substances that are volatile	B	eExam
<input type="checkbox"/>	MCQ	The product obtained from gravimetric analysis is of limited _____.	Solubility	Quantity	None of the above	All of the above	A	eExam
<input type="checkbox"/>	MCQ	Gravimetry is a _____ method of analysis.	Qualitative	Quantitative	Qualitative and quantitative	Spectroscopic	B	eExam
<input type="checkbox"/>	MCQ	The component of a spectrometer that converts radiant energy to electrical energy is called _____.	Detector	Read-out	Source	Monochromator	A	eExam
<input type="checkbox"/>	MCQ	The absorption of radiation can be used either for _____.	Qualitative analysis	Qualitative or quantitative analysis	Quantitative analysis	Detection of atoms or molecules	B	eExam

<input type="checkbox"/>								
<input type="checkbox"/>	MCQ	The second step of the process of absorption of electromagnetic radiation by some species is called _____.	Absorption of radiation	Atomic or molecular transition	Relaxation	Transit state	C	eExam
<input type="checkbox"/>	MCQ	The absorption of electromagnetic radiation by some species is considered to undergo a _____ process.	Four-step	Three-step	Five-step	Two -step	D	eExam
<input type="checkbox"/>	MCQ	The molecules at ordinary room temperature is considered to be at _____.	Excited state	Transit state	Lowest electronic energy state	Atomic or molecular state	C	eExam
<input type="checkbox"/>	MCQ	There are how many basic processes by which molecules can absorb radiation?	Two	Three	Four	One	B	eExam
<input type="checkbox"/>	MCQ	_____ is converting the accessible sample matrices to the form which is accessible to instrument.	Sampling	Mixing solids laboratory samples	Dissolution	Sample decomposition	D	eExam
<input type="checkbox"/>	MCQ	_____ is a method of sample selection which involves pouring the sample so that it takes on a conical shape, flattening the conical shape out into a cake, the cake is then divided into quarters and two quarters which face opposite one another are discarded whilst the other two are combined and constitute the reduced sample.	Quartering	Coning	Quartering and coning	Flattening and quartering	C	eExam
<input type="checkbox"/>	MCQ	Grinding and crushing of a sample does which of the following?	Reduce the size of particles of sample materials and increase the surface area.	Mix the samples in order to ensure random distribution of the components in the analytical sample.	Convert sample material in solid form to solution.	Reduce the water content of a sample material.	A	eExam
<input type="checkbox"/>	MCQ	In gravimetric analysis, the sample to be determined is _____.	Extracted with an organic solvent to obtain a precipitate	Chemically reacted with a reagent to yield a precipitate	Dissolved in an aqueous solvent and filtered to obtain the precipitate	Extracted with an inorganic solvent to obtain the precipitate	B	eExam
<input type="checkbox"/>	MCQ	The _____ can be arbitrarily broken down into different region according to wavelength.	Frequency	Electromagnetic spectrum	Photon	Wave number	B	eExam

<input type="checkbox"/>	MCQ	_____ is a branch of analytical chemistry that deals with the study of concentration of analyte as a function of amount of radiation absorbed when electromagnetic radiation from appropriate source is directed at it.	Titrimetry	Spectroscopy	Redox titration	Complexometry	B	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	The number of _____ that complexes metals depends on the coordination number of the metal.	Non metals	electrodes	ions	Ligands	D	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	_____ are complexing agents that bind with metals to form complexes.	Ligands	Electrodes	Salt	Electrolytes	A	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	Which of the following is not among the chain of custody procedure?	Sample label	Sample collection	Assignment of sample for analysis	Sample delivery to the laboratory	B	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	Which of these statements is true?	Precision helps in expressing confidence interval	Significant figure is a method of showing dispersion from central values.	The smaller the standard deviation, the more closely the data are clustered.	Average deviation is a method of expressing the minimum number of digit of a data.	C	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	_____ enables analytical chemist to accept conclusions that have high probability of being correct and to reject conclusions that are doubtful.	Performance of replicate measurements	Statistical analytical tools	None of the above	All of the above	B	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	_____ are single samples collected at specific spot at a site over a short period of time.	Composite samples	Integrated samples	Grab samples	Non composite samples	C	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	_____ is designed to investigate whether there is a significant difference between two method based on their standard deviation.	Correlation	Coefficient of variation	F-test	Variance	C	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	_____ cannot be predicted or estimated.	Random error	Determinate error	Systematic error	Controllable error	A	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	_____ is a form of volumetric analysis which involves the transfer of electrons from one substance to another.	Precipitation titration	Complexometric titration	Acid-base titration	Oxidation-reduction titration	D	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	The magnitude of random errors can be reduced by _____.	Using a well calibrated equipment	Replication of experiment	Exercising carefulness during the experiment	Using different analytical methods to perform the same experiment	B	<input type="checkbox"/> eExam

<input type="checkbox"/>	MCQ	Co-precipitation with impurities, incomplete reaction, impurities in the reagents used is an example of _____.	Operative error	Methodic error	Instrumental error	Random error	B	eExam
<input type="checkbox"/>	MCQ	Characteristics of a standard solution include _____.	A method must exist for detecting the equivalence point between the standard solution and analyte.	The reaction of the standard solution with the analyte must be describable by equation.	Reaction of the standard solution with analyte should be rapid.	All of the above.	D	eExam
<input type="checkbox"/>	MCQ	_____ are also known as indeterminate errors.	Random error	Systematic error	Uncontrollable error	Controllable error	A	eExam
<input type="checkbox"/>	MCQ	Which of the following is true?	Treatment or preparation of sample converts the sample from the nature in which it exists at the site of sampling to the form in which it can be analysed.	The chain of custody procedure is not helpful in identifying the source of contamination of a sample.	All of the above	None of the above	A	eExam
<input type="checkbox"/>	MCQ	The titrand is _____.	The reagent added from burette to the analyte in a conical flask.	The solvent used in making a solution of the base.	The solvent used in diluting the acid.	The analyte in a conical flask to which reagent in the burette is added to.	D	eExam
<input type="checkbox"/>	MCQ	Completion of a titration reaction can be detected by _____.	Increase in the volume of the analyte in the conical flask.	Decrease in the volume of the reagent in the burette	Observation of marked change of pH in the titration of an acid with a base.	All of the above.	C	eExam
<input type="checkbox"/>	MCQ	Using an incorrectly standardised pH meter in an experiment will introduce _____.	Random error	Variation in measurement	Systematic error	Uncontrollable error	C	eExam
<input type="checkbox"/>	MCQ	Which of these statements is not true?	In titration reaction, there should be no side or parallel reaction.	In titration reaction, there should be known reaction pattern between the analyte and titrant.	Titration reaction need not be rapid.	Titration reaction should be quantitative.	C	eExam
<input type="checkbox"/>	MCQ	The titrant is _____.	The reagent added from burette to the analyte in a conical flask.	The solvent used in making a solution of the base.	The solvent used in diluting the acid.	The analyte in a conical flask to which reagent in the burette is added to.	A	eExam
<input type="checkbox"/>	MCQ	The likelihood that the true value falls within a range is called _____.	Confidence limit	Confidence level	Confidence interval	Accuracy	B	eExam

<input type="checkbox"/>	MCQ	Write 92500 as 3 significant figures.	9.25×10^4	9.250×10^4	925	0.925	A	eExam
<input type="checkbox"/>	MCQ	Which of the following is not a method used for the prediction and correction of systematic error.	Analysing samples of unknown composition	Analysing blank samples	Using different analytical methods to measure the same quantity	By operators of varying capabilities in different laboratories carrying out the same analyses.	A	eExam
<input type="checkbox"/>	MCQ	A miniature replica of the bulk of materials to be analysed is called _____.	Grab sample	Composite sample	Gross sample	Integrated sample	C	eExam
<input type="checkbox"/>	MCQ	A good _____ must have the same characteristic or features with that of the original population from where it is selected.	Raw material	Sample	Analysis	Result	B	eExam
<input type="checkbox"/>	MCQ	The range within which the true value might fall within a given probability is known as?	Coefficient of variation	Accuracy	Confidence limit	Confidence interval	D	eExam

Showing 1 to 120 of 120 entries

Previous 1 Next