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Question Type	Question	A	B	C	D	Answer	Remark
FBQ	The amount of substance dissolved in a certain amount of solvent is referred to as _____.	Solubility					eExam
FBQ	The type of tautomerism which involves a shift in interatomic distance within a molecule, without the separation of any atom from the rest of the molecule , as an intermediate stage is called _____.	Valence tautomerism					eExam
FBQ	Solubility of branched isomers is _____ than that of straight chain isomers	more					eExam
FBQ	Branched chain hydrocarbons have lower _____ than their straight chain isomers.	Melting point	Boiling point				eExam
FBQ	Unbranched isomer has higher _____ than branched isomer	boiling point					eExam
FBQ	The extent to which acids transfer a proton to a standard base determines _____.	Strength of acid					eExam
FBQ	The definition of an acid as a proton donor was given by _____.	Bronsted and Lowry					eExam

☐								
☐	FBQ	Two factors that affect the boiling point of a substance are _____ and _____.	Polarity,Molecular structure and London forces					eExam
☐	FBQ	A change in molecular structure which affect the reactivity of the molecule by changing the electron distribution is called _____	Electronic effect					eExam
☐	FBQ	The stronger the acid, the smaller the _____.	pKa					eExam
☐	FBQ	Hydrogen no - bond resonance is otherwise called _____.	Hyper conjugation					eExam
☐	FBQ	The influence of the inductive effect on acid strength is almost negligible after the _____ carbon in the chain.	Fourth					eExam
☐	FBQ	The method of reduction of the carbonyl group of ketone produced by Friedel - craft is _____.	Clemmenson reduction					eExam
☐	FBQ	The reaction of aromatic compounds with alkyl halide in the presence of anhydrous $$AlCl_{3}$$ as catalyst is called _____.	Friedel - Craft alkylation					eExam
☐	FBQ	The backbone of an organic molecule is _____.	Carbon - carbon bond	C - C bond				eExam
☐	FBQ	Cyclohexane has _____ fewer hydrogens than n-hexane.	2					eExam

☐	FBQ	The rule that states that in a dehydrohalogenation reaction of alkyl halides, the major product will be the product that has the more alkyl groups attached to the resultant carbon - carbon double bond is _____ rule.	Saytzeff					eExam
☐	FBQ	E2 elimination reaction can also be referred to as _____ reaction.	Stereoselective anti-elimination					eExam
☐	FBQ	The dehydrohalogenation of alkyl halides yields _____.	Alkenes					eExam
☐	FBQ	A side reaction that occurs during substitution reactions of alkyl halides is _____	Dehydrohalogenation	Elimination of hydrogen halide				eExam
☐	FBQ	The reaction in which aryl halides are formed from aromatic amines is called _____.	Sandmeyer reacion					eExam
☐	FBQ	Oxidation of naphthalene under controlled condition yields _____	14-naphthoquinone					eExam
☐	FBQ	Fused benzenoid hydrocarbons with two benzene rings is _____.	Naphthalene					eExam
☐	FBQ	Fused benzenoid hydrocarbon with three benzene rings is _____ .	Anthracene	Phenanthrene				eExam
☐	FBQ	In keto-enol tautomers, the tautomer that predominates is _____.	Keto tautomer					eExam
☐	FBQ	Tautomers which differ from each other only in the location of a hydrogen atom and a double bond are called _____.	Proton tautomers					eExam

☐	FBQ	In keto-enol tautomers, the tautomer that is more stable is _____.	Keto tautomer					eExam
☐	FBQ	In terms of aromatization, naphthalene has _____ π electrons	10					eExam
☐	FBQ	Naphthalene has _____ number of resonance hybrid structures	Three					eExam
☐	FBQ	Reduction of benzene using Nickel at 425-525K and 25atm yields _____.	Cyclohexane					eExam
☐	FBQ	The condition for reduction of benzene is _____ or _____	Nickel 425-525K 25atm, sodium in liquid ammonia in the presence of ethanol.					eExam
☐	FBQ	Reduction of benzene is by _____ of benzene.	Hydrogenation					eExam
☐	FBQ	The reagent for nitration of benzene is _____.	mixture of conc. Trioxonitrate v acid and tetraoxosulphate vi acid					eExam
☐	FBQ	The major product in the nitration of nitrobenzene is _____	meta-dinitrobenzene					eExam
☐	FBQ	Nitration of nitrobenzene yields _____ .	meta - dinitrobenzene	Para - dinitrobenzene				eExam
☐	FBQ	Isomers which are related to each other with the actual movement of electrons as well as one or more atoms and can reversibly be interconverted are called _____	Tautomers					eExam
☐	FBQ	In terms of steric effect, _____ conformational isomer is more stable.	Staggered					eExam
☐	FBQ	In terms of steric effect _____ geometric isomer is more stable.	Trans					eExam

☐	FBQ	Types of conformational isomers are _____ and _____.	Staggered, Eclipsed					eExam
☐	FBQ	The types of geometric isomers are trans and _____.	Cis-isomer					eExam
☐	FBQ	The effect arising from the spatial interactions between the groups in a compound is called _____.	Steric effect					eExam
☐	FBQ	Hyperconjugation is also known as _____	Hydrogen no-bound resonance					eExam
☐	FBQ	Inductive effect is a type of _____ effect	electronic					eExam
☐	FBQ	Chloro compounds can be used in manufacturing of _____.	DDT					eExam
☐	FBQ	The reaction of ethanal with phosphorus pentachloride yields _____.	11-dichloroethane					eExam
☐	FBQ	Tertiary alkyl halides undergo substitution by _____ mechanism	SN1					eExam
☐	FBQ	On the basis of the mechanism of substitution reactions, nucleophilic substitution reaction can be divided _____.	Two.					eExam
☐	FBQ	The type of dihalides in which the two halogens atoms are attached to adjacent carbon atoms is called _____.	Vicinal dihalides					eExam
☐	FBQ	Chlorination of ethene yields _____.	Chloroethene					eExam

☐								
☐	FBQ	The type of dihalides in which both halogens are attached to the same carbon atoms is called _____.	Geminal dihalides					eExam
☐	FBQ	The halogen derivatives of aromatic compounds are _____	Aryl halides					eExam
☐	FBQ	The reaction between benzene and acetyl chloride (acid halide) in the presence of aluminium trichloride yields _____.	Aceto[phenone					eExam
☐	FBQ	Ortho, Para and meta-direction are effects of _____.	Effect of substituents on orientation .					eExam
☐	FBQ	Molecular structure, London forces, Polarity of a compound and Hydrogen bonding are factors that affect _____	Boiling point					eExam
☐	FBQ	An important intermolecular force _____	Dipole-dipole interactions or London forces	Hydrogen bonding				eExam
☐	FBQ	The actual molecule or ion of a covalent molecule with more than one lewis structure is called _____.	Resonance hybrid					eExam
☐	FBQ	The lewis structure of a covalent molecule are otherwise called _____.	Resonance structure	Resonance contributors				eExam
☐	FBQ	The phenomenon of the transmission of charge through a chain of atoms linked by δ-δ bonds is called _____.	Inductive effect					eExam
☐	FBQ	A change in molecular structure can occur as a result of _____.	Steric effect	Electronic effect				eExam

	FBQ	The acidic and basic properties of molecules are influenced by _____ and _____	Functional groups, Structural variations					eExam
	MCQ	The symbol of resonance is ----------.	$$< —>$$	$$<¥>$$	$$—>$$	$$<—$$	B	eExam
	MCQ	The process by which benzene and its derivatives are extracted from petroleum is called -----------	Polymerization	Hydroforming	Cracking	Aromatization	B	eExam
	MCQ	The presence of an aromatic ring in a compound is detected by ----------.	Mass spectroscopy	Infrared spectroscopy	Ultraviolet spectroscopy	Chromatography	C	eExam
	MCQ	The cannonical forms of benzene were proposed by ?	Kekule and Dalton	Dalton and Dewar	Kekule and Dewar	None of the above	C	eExam
	MCQ	The decreasing order of basicities is:	$$NH_{3}< CH_{3}NH_{2}< NH_{2}OH$$	$$NH_{3}> CH_{3}NH_{2}> NH_{2}OH$$	$$CH_{2}NH_{2}> NH_{3} >NH_{2}OH$$	$$CH_{3}NH_{2} <NH_{3} < NH_{2}OH$$	D	eExam
	MCQ	The substitution of an acyl group into an aromatic ring by the reaction with an acid chloride in the presence of Lewis acid is called ----------.	Friedel - Craft acylation	Friedel - Craft akylation	Hyperconjugation	Hydrogenation	A	eExam
	MCQ	There are ---------- resonance structures of benzene	Two	Three	Four	Five	D	eExam
	MCQ	The method of reduction of the carbonyl group of ketone produced by Friedel -crafts acylation is ---------- reduction.	Electrophilic	Acylation	Alkylation	Clemmenson	D	eExam
	MCQ	Sulphonation of benzene yields ----------.	Sulphuric acid	Sulphonic acid	auto-sulphur benzene	Para-sulphur benzene	B	eExam
	MCQ	The reaction of aromatic compounds with alkyl halide in the presence of anhydrous $$ $AlCl_{3}$$ as catalyst is called ----------	Friedel - Craft acylation	Friedel - Craft akylation	Hyperconjugation	Hydrogenation	B	eExam
	MCQ	The number or π-electrons in benzene is?	2	4	6	8	C	eExam
	MCQ	The stronger an acid the ---------- is its conjugate base	Stronger	Weaker	Both have no correlation	None of the above	B	eExam
	MCQ	Down the group, electronegativity ----------.	Increases	Decreases	Neutral	None of the above	B	eExam

	MCQ	The parent compound of polynuclear hydrocarbons is ---------.	Naphthalene	Anthracene	Phenanthrene	Biphenyl	A	eExam
	MCQ	A benzenoid hydrocarbon in which two benzene rings are fused together at the ortho position is called ----------.<	Naphthalene	Anthracene	Phenanthrene	Biphenyl	A	eExam
	MCQ	The benzenoid hydrocarbons in which two or more benzene rings are fused together at ortho position in such a way that each pair of rings shares rings shares two carbons are called ---------- benzenoid hydrocarbon.	Isolated	Condensed	1 and 2	None of the above	B	eExam
	MCQ	An example of Isolated Benzenoid Hydrocarbon is ----------.	Naphthalene	Phenanthrene	Biphenyl	Anthracene	C	eExam
	MCQ	The catalyst usually employed in Friedel-Craft acylation is -------.	Tetraoxosulphate VI acid	Carbon tetrachloride	Sodium hydroxide	Aluminium trichloride	D	eExam
	MCQ	The preparation of aryl halides which involves the aromatic ammine first being converted to diazonium salt and then to aryl halide uisng solution of cuprous halide dissolved in concentrated halogen acid is called ---------.	Friedel-Crafts Acylation	Friedel -Crafts Alkylation reaction	Hunsdiecker reaction	Sandmeyer reaction	D	eExam
	MCQ	In electrophilic addition to alkenes, the electrophilic attack is on ---------.	H-H bond	C-H bond	C=C bond	None of the above	C	eExam
	MCQ	Another name for methyl benzene is _____.	Xylene	Phenol	Toluene	Naphthalene	C	eExam
	MCQ	In the hydrogenation of benzene,the condition for reaction is _____.	Ni 525-625K 25atm	Ni 425-525K 25atm	Fe 525-625K 25atm	Fe 425-525K 25atm	B	eExam
	MCQ	Hydrogenation of benzene at higher temperature and under pressure yields -------	Cyclohexane	Cyclohexene	Cyclohexyne	Benzaldehyde	A	eExam
	MCQ	All these are electron donating groups with respect to reactivity and orientation EXCEPT _____	$$-OH$$	$$-OR$$	$$-NHCOR$$	$$-CN$$	D	eExam
	MCQ	The E2 reactions of alkyl halides are favoured by the use of ---------.	Strong acid	Strong base	Weak acid	Weak base	B	eExam
	MCQ	In the preparation of trichloromethane by chlorination of methane in the presence of nitrogen, the catalyst required is ----------.	Sodium chloride	Potassium chloride	Hydrogen chloride	Cupric chloride	D	eExam

	MCQ	Bleaching powder provides chlorine which oxidizes ethanol to _____.	Ethane	Ethene	Ethanal	Ethanone	C	eExam
	MCQ	The reaction of grignard reagent and aldehyde other than methanal followed by decomposition of the addition product with a dilute acid gives ----------.	Primary alcohols	Secondary alcohols	Tertiary alcohols	Ketones	B	eExam
	MCQ	For halogen derivatives of aromatic compounds, boiling point ---------- with increase in atomic size of the halogen atom.	Decreases	Increases	Averages	Has no effect	B	eExam
	MCQ	The carbon-halogen sigma bond in alkyl halide results through overlap of the ------- hybrid orbital with the p orbital of the halogen atom.	$$sp$$	$$sp^{2}$$	$$sp^{3}$$	None of the above	A	eExam
	MCQ	Chlorination of ethene at high temperature yields ---------.	Chloroethene	12-dichloroethene	13-dichloroethene	Chloroethyne	A	eExam
	MCQ	The compounds formed when hydrogen halides react with alkenes are ---------.	Alkyl halides	Alkenyl halides	Vinylic halides	Vinyl halides	A	eExam
	MCQ	The most widely used method for the preparation of alkyl halides is from _____.	Alkenes	Alkanes	Grignard reagent	Alcohols	D	eExam
	MCQ	The halogen derivative of aromatic compounds in which two halogen atoms are attached to adjacent carbon atom is called _____.	Geminal dihalide	Vicinal dihalide	Alkenyl dihalide	Alkyl dihalide	B	eExam
	MCQ	The halogen derivative of aromatic compounds in which two halogen atoms are attached to the same carbon atom is called ---------.	Geminal dihalide	Vicinal dihalide	Alkenyl dihalide	Alkyl dihalide	A	eExam
	MCQ	The energy difference between the expected value of heat evolved when hydrogen is added to benzene and the experimental heat value is called ----------.	Ionization energy	Resonance energy	Stabilization energy	Potentential energy	B	eExam
	MCQ	A compound in which a halogen atom is attached to another carbon atom by a double bond is called -----.	Alkyl halide	Vinyl halide	Geminal dihalides	Vicinal dihalides	B	eExam

	MCQ	The number of classes to which halogen derivatives of aromatic compounds could be classified based on the nature of the hydrocarbon residue to which the halogen atom is attached is----------.	Two	Three	Four	five	B	eExam
	MCQ	The expression relating pKa and pH is known as ---------- equation.	Henderson-Hasselbalch	Bronsted-Lowry	Arrhenium	None of the above	A	eExam
	MCQ	The molecular reactivity of a substance depends on ---------.	Number of carbon atoms in the substance	Functional groups present in the substance	Number of hydrogen present in the substance	type of bonds in the substance	C	eExam
	MCQ	Induced dipole - induced dipole interactions are also called ---------.	Dipole -dipole interactions	Hydrogen bonding	London forces	Van der Waal forces	A	eExam
	MCQ	Increase in melting point as a result of increase in molecular weight is due to ---------.	Hydrogen bonding	Dipole -dipole interactions	Intramolecular forces	London forces	D	eExam
	MCQ	Pure crystalline solids have ---------- melting points.	Low	High	Moderate	have no effect on	B	eExam
	MCQ	Dipole - dipole interactions and induced dipole - induced dipole interactions are collectively called ----------.	Intramolecular forces	Intermolecular forces	London forces	Van der Waal forces	D	eExam
	MCQ	A substance which ionises to produce hydroxide ions is called ----------.	Acid	Base	Salt	None of the above	B	eExam
	MCQ	A benzene ring is a ---------- system.	Nonconjugation	Conjugation	Hyperconjugation	None	B	eExam
	MCQ	With respect to solubility, the more soluble isomer is the---------- isomer.	Branched	Straight chain	Condensed	Optical	A	eExam
	MCQ	Increase in intermolecular forces results in ---------- solubility.	High	Low	Average	Has no effect	B	eExam
	MCQ	The process by which a substance dissolves in a solvent and its constituent ions get separated from each other is called ----------.	Solubility	Salinity	Dissolution	Solvation	D	eExam
	MCQ	The properties of a solvent which measures its ability to separate the ions of the solute is --------.	Density	Dielectric constants	Solubility	Melting point	B	eExam
	MCQ	All these affect boiling points except ----------.	Molecular structure	London forces	Polarity	Shape of molecule	D	eExam
	MCQ	A chief source of benzene and its derivatives is ----------.	Coaltar	Bitumen	Limestone	Basalt	A	eExam

☐	MCQ	The point at which the vapour pressure of a liquid is equal to external pressure is called ----------.	Melting point	Boiling point	Equivalence point	Critical point	B	eExam
☐	MCQ	The heating of toluene and hydrogen under pressure in the presence of a metal catalyst yields ----------.	Pyridine	Xylene	Naphthalene	Benzene	D	eExam
☐	MCQ	Geometric isomerism yields ---------- isomers	Two	Three	Four	Five	A	eExam
☐	MCQ	The geometric isomer that has higher melting point is -------- isomer.	cis	Trans	Dextro	Laevo	B	eExam
☐	MCQ	The interactions that exist between non polar molecules is called ----------.	London forces	Hydrogen bonding	dipole-dipole interactions	Van der Waals forces	A	eExam
☐	MCQ	London forces, hydrogen bonding, and dipole-dipole interactions are ---------.	Covalent forces	Molecular forces	Intermolecular forces	Intramolecular forces	C	eExam
☐	MCQ	The (4n + 2)pie-electron is -------- rule .	Saytzeff's	Hoffman's	Huckel's	Hund's	C	eExam
☐	MCQ	The simplest aromatic compound is called ----------.	Alkanones	Naphthalene	Benzene	Phenol	C	eExam

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