

eExam Question Bank

Coursecode:

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<input type="checkbox"/>	Question Type	Question	A	B	C	D	Answer	Remark
<input type="checkbox"/>	FBQ	<input type="text"/> are formed when vapourised carbon condenses in an atmosphere of inert gas and produces clusters with 60 carbon atoms or 70	fullerenes	fullerenes				<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	<input type="text"/> is the only liquid halogen.	Bromine	Br				<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	The only halogen that can oxidize chlorine is <input type="text"/>	fluorine	F				<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	Interhalogens have physical properties that are intermediate between those of their parent halogens, <input type="text"/> (True/false)	1	1				<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	If X and Y are two different halogens, then the X-Y bond is more polar than the X-X or Y-Y bond. <input type="text"/> True/false	1	1				<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	Carbon is more electropositive than silicon, <input type="text"/> (True/false.)						<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	The element known as 'super halogen' is <input type="text"/>	fluorine					<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	Of all halogens, <input type="text"/> has the highest ionization energy	fluorine					<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	<input type="text"/> has the highest electronegativity value	fluorine					<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	The halogen <input type="text"/> is solid at room temperature	iodine					<input type="button" value="eExam"/>

<input type="checkbox"/>									
<input type="checkbox"/>	FBQ	The halogen which has only two complete electron shells below the valence shell is <input type="text"/>	chlorine						eExam
<input type="checkbox"/>	FBQ	The oxide of nitrogen which is colourless and rather unreactive is <input type="text"/>	Nitrogen(I)oxide	nitrous oxide					eExam
<input type="checkbox"/>	FBQ	The ground state electronic configuration of phosphorus is <input type="text"/>	1s2 2s2 sp6 3s2 3p3						eExam
<input type="checkbox"/>	FBQ	Nitriles are oxidized to nitrates by <input type="text"/>	nitrifying	nitrifying					eExam
<input type="checkbox"/>	FBQ	Ammonium compounds are converted into nitriles by <input type="text"/> bacteria	nitrosifying	nitrosifying					eExam
<input type="checkbox"/>	FBQ	Bacteria which convert nitrogen compounds into free nitrogen are known as <input type="text"/> bacteria	denitrifying	denitrifying					eExam
<input type="checkbox"/>	FBQ	Bacteria which convert nitrogen into its compounds are called <input type="text"/> bacteria	nitrogen-fixing	nitrogen fixing					eExam
<input type="checkbox"/>	FBQ	Conversion of nitrogen into its compounds is known as <input type="text"/>	Nitrogen fixation	nitrogen fixation					eExam
<input type="checkbox"/>	FBQ	Continuous exchange of nitrogen between the atmosphere and the biosphere is called <input type="text"/>	nitrogen cycle	nitrogen cycle					eExam
<input type="checkbox"/>	FBQ	Hydrazine is used as a <input type="text"/> along with liquid air or oxygen as oxidant.	rocket fuel	rocket fuel					eExam
<input type="checkbox"/>	FBQ	The molecular formula of hydrazine is <input type="text"/>	N ₂ H ₄						eExam
<input type="checkbox"/>	FBQ	Ammonia acts as a Lewis base because it possesses a <input type="text"/>	lone pair of electrons	lone electrons					eExam
<input type="checkbox"/>	FBQ	The hybridization of nitrogen in ammonia is <input type="text"/>	sp ³						eExam

<input type="checkbox"/>									
<input type="checkbox"/>	FBQ	All group V elements except <input type="text"/> react with sulphur to form sulphides.	Nitrogen	N					eExam
<input type="checkbox"/>	FBQ	Red phosphorus is more reactive than white phosphorus; <input type="text"/> (True/false)							eExam
<input type="checkbox"/>	FBQ	The allotrope of phosphorus used in matches and fireworks is <input type="text"/>	red phosphorus	red, phosphorus red					eExam
<input type="checkbox"/>	FBQ	<input type="text"/> are the fibrous silicates, having high tensile strength and heat/fire resistance, used as insulators and steam pipe lagging.	Asbestos	asbestos					eExam
<input type="checkbox"/>	FBQ	The 'great balls of carbon' in which carbon atoms are symmetrically arranged in closed shells are known as <input type="text"/>	Fullerenes	fullerene					eExam
<input type="checkbox"/>	FBQ	Of the group IV element, the one that exhibits anomalous behaviour is <input type="text"/>	Carbon	C					eExam
<input type="checkbox"/>	FBQ	Zeolites are also called <input type="text"/>	Permutits	Permutits					eExam
<input type="checkbox"/>	FBQ	<input type="text"/> is obtained when of silicates mainly of sodium and calcium is melted and supercooled	glass	transparent solid					eExam
<input type="checkbox"/>	FBQ	Polytetrafluoroethane (PTFE) is commercially known as <input type="text"/>	Teflon	teflon					eExam
<input type="checkbox"/>	FBQ	<input type="text"/> is a property by virtue of which elements form long chain compounds by single or multiple bonds between atoms of the same element.	Catenation	catenation					eExam
<input type="checkbox"/>	FBQ	The form of carbon that conducts electricity is <input type="text"/>	graphite	graphite					eExam

<input type="checkbox"/>									
<input type="checkbox"/>	FBQ	The group of elements in the periodic table that is associated with semiconductor properties is <input type="text"/>	Four	IV					eExam
<input type="checkbox"/>	FBQ	Silicon and germanium do not exhibit allotropy <input type="text"/> (True/false)	1	1					eExam
<input type="checkbox"/>	FBQ	The second most abundant element in the earth's crust after oxygen is <input type="text"/>	Silicon	Silicon(IV) oxide					eExam
<input type="checkbox"/>	FBQ	The only group IV element that occurs in the elemental state is <input type="text"/>	Carbon	C					eExam
<input type="checkbox"/>	FBQ	The most metallic element in group IV is <input type="text"/>	Lead	Pb					eExam
<input type="checkbox"/>	FBQ	Elements having some metallic and nonmetallic properties are called <input type="text"/>	metalloids	metalloid					eExam
<input type="checkbox"/>	FBQ	A derivative of <input type="text"/> is used in making photographic plates	bromine	Br					eExam
<input type="checkbox"/>	FBQ	Teflon is formed from a derivative of <input type="text"/>	fluorine	F					eExam
<input type="checkbox"/>	FBQ	The thermal stability of hydrogen halides decreases from HF to HI <input type="text"/> (True/false)	1	1					eExam
<input type="checkbox"/>	FBQ	The strongest acid among the hydrogen halides is <input type="text"/>	HI	hydroiodic acid					eExam
<input type="checkbox"/>	FBQ	Halogen means <input type="text"/>	salt producer	salt former					eExam
<input type="checkbox"/>	FBQ	<input type="text"/> is a gas formed by the action of an acid on a carbonate	CO ₂	Carbondioxide					eExam
<input type="checkbox"/>	FBQ	Zeolites are used in water as <input type="text"/>	softeners	softener					eExam

<input type="checkbox"/>								
<input type="checkbox"/>	FBQ	The group V metal that can not be obtained from its sulphide ore is <input type="text"/>	Nitrogen	N				eExam
<input type="checkbox"/>	FBQ	Mixed chloroflorocarbons are known as <input type="text"/>	freons	freon				eExam
<input type="checkbox"/>	FBQ	Arsenic has a <input type="text"/> odour	garlic-like	garlic				eExam
<input type="checkbox"/>	FBQ	Silicones are a group of <input type="text"/>	Organosilicon polymers	organosilicon				eExam
<input type="checkbox"/>	FBQ	The purest form of silicon dioxide is <input type="text"/>	rock crystal	rock crystal				eExam
<input type="checkbox"/>	FBQ	The group V element that forms alloys with metals is <input type="text"/>	Bismuth	Bi				eExam
<input type="checkbox"/>	FBQ	The most reactive form of phosphorous is <input type="text"/>	brown phosphorus	brown				eExam
<input type="checkbox"/>	FBQ	The oxide of nitrogen that gives a positive brown ring test is <input type="text"/>	nitric oxide	nitrogen(II) oxide				eExam
<input type="checkbox"/>	FBQ	High purity silicon is obtained from <input type="text"/>	Silicon tetrachloride	Silicon(IV) oxide				eExam
<input type="checkbox"/>	FBQ	An example of amorphous form of carbon is <input type="text"/>	lamp black	soot				eExam
<input type="checkbox"/>	FBQ	The structure of graphite is <input type="text"/>	hexagonal	planar hexagonal				eExam
<input type="checkbox"/>	FBQ	The two crystalline allotropies of carbon are <input type="text"/>	Diamond and graphite	diamond & graphite				eExam
<input type="checkbox"/>	FBQ	Clays are essentially alumina-silicates of <input type="text"/>	sodium	Calcium				eExam
<input type="checkbox"/>	MCQ	Which of these is not a carbide?_____	dative	ionic	covalent	interstitial	A	eExam
<input type="checkbox"/>	MCQ	When carbon is linked to elements of lower or approximately the same electronegativity, the binary compounds formed are termed	hydroxides	oxides	halides	carbides	D	eExam
<input type="checkbox"/>	MCQ	Solid Carbondioxide is known as_____	hydrocarbon	Teflon	dry ice	carbene	C	eExam

<input type="checkbox"/>	MCQ	Silicates are salts of _____	Silica gel	silicon tetraoxide	tetrachloro silicon	Crystaamite	B	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	Lead dioxide is used as _____	bleach	dentoator	gun powder	anti knock	B	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	Zeolites is used in water as _____	oxidants	purifiers	disinfectants	softener	D	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	Which of the group V elements give only one oxidation state?	bismouth	nitrogen	lead	phosphorus	A	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	Which of the group V oxoacids is manufactured by Ostwald process?	nitric acid	dinitrogen tetraoxide	nitrogen hydroxide	nitrous oxide	A	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	Aqua regia is a mixture of Nitric acid and _____	limewater	concentrated Hydrochloric acid	water	conc sulphuric acid	B	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	Which of the group V metals can not be obtained from its sulphide ore?	lead	Bismouth	Nitrogen	phosphorus	C	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	Which of these is not an interstitial carbide?	tungsten carbide	Vanadium carbide	boron carbide	titanium carbide	C	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	Which of these is true of interstitial carbide?	They are binary compounds	they are amorphous compounds	they are soluble compounds	they are non stoichiometric compound	D	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	The following are crystaline forms of silica, except----- ----	quartz	Crystaamite	tridymite	cristobalite	B	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	The heavier elements of group IV functions as-----	lewis acid	lewis base	Lowry base	Lowry acid	A	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	Which of these is an example of covalent carbides?	dicloro silicon	SiC	tetrachloro silicon	TiC	B	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	Mixed chloroflorocarbons are known as -----	halides	carbenes	teflon	Freons	D	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	Which of the following is responsible for the easy hydrolysis of lead halides ?	Availability of d-orbitals and low bond energy of Pb-H bond	availability of lone pair of electron in P orbital	availability of lone pair of electron in d orbital	completely filled d orbitals by electron	A	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	One of these is not a form of silica	Flint	tridymite	cristobalite	Freons	D	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	The purest form of silicon dioxide is----	cristobalite	silibolite	rock crystal	silicates	C	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	Which of these is not a raw material for the cement manufacture?	alumina	Silicate	lime	silica	B	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	Which of these is not a component of lead-acid battery?	Tetraoxosulphate(VI) acid	lead oxide	lead chloride	spongy lead	C	<input type="button" value="eExam"/>

<input type="checkbox"/>								
<input type="checkbox"/>	MCQ	Which of these statements is not true of group V elements ?	Phosphorus exists both in gaseous and solid forms	Nitrogen exists as a diatomic molecule	bismuth is predominantly non- metallic in nature	antimony is a metalloid	C	eExam
<input type="checkbox"/>	MCQ	Nitrogen is used for all but one of the following	freezing of foodstuffs	freeze grinding of normally soft and rubbery materials	in making fire works	in the packaging of processed foods	C	eExam
<input type="checkbox"/>	MCQ	Nitrogen is obtained commercially from	rock	air	coke	water vapour	B	eExam
<input type="checkbox"/>	MCQ	Which of these is not a group V element?	Nitrogen	Arsenic	Phosphorus	Calcium	D	eExam
<input type="checkbox"/>	MCQ	A derivative of ----- is used in making photographic plates	Chlorine	Bromine	Fluorine	Iodine	B	eExam
<input type="checkbox"/>	MCQ	PVC pipes are manufactured using	Chlorine	Bromine	Fluorine	Iodine	A	eExam
<input type="checkbox"/>	MCQ	Teflon is formed from a derivative of-----	Chlorine	Bromine	Fluorine	Iodine	C	eExam
<input type="checkbox"/>	MCQ	Deficiency of one of these leads to goitre in humans	Chlorine	Bromine	Fluorine	Iodine	D	eExam
<input type="checkbox"/>	MCQ	Which of these halogens shows no basic properties	Chlorine	Bromine	Fluorine	Iodine	C	eExam
<input type="checkbox"/>	MCQ	The thermal stability of hydrogen halides-----	increases from HF to HI	decreases from HF to HI	is constant across the group	has no fixed pattern across the group	A	eExam
<input type="checkbox"/>	MCQ	Which of these halogen oxides is useful in the estimation of CO	Chlorine	Bromine	Fluorine	Iodine	D	eExam
<input type="checkbox"/>	MCQ	One of these statements is not true of oxides of halogen.	All oxides of halogens have positive free energies of formation	The higher oxides tend to be more stable than the lower	Except for that of iodine all oxides tend to be explosive	Chlorine oxides are used as bleaching agents	A	eExam
<input type="checkbox"/>	MCQ	Which of these halogens does not form oxoacids ?	Chlorine	Bromine	Fluorine	Iodine	C	eExam
<input type="checkbox"/>	MCQ	Compounds formed by the interaction of one halogen with other halogens are called -----	pseudo halogens	interhalogens	Polyhalides	polyhalogens	B	eExam
<input type="checkbox"/>	MCQ	One of these is not a pseudo halogen.	cyanogen	thiocyanogen	selenocyanogen	Teflon	D	eExam
<input type="checkbox"/>	MCQ	Fluorine reacts with alkalis to yield-----	the salt	the hypohalite	the oxide	the acid	C	eExam
<input type="checkbox"/>	MCQ	Halogen means-----	salt producer	reducing agent	oxidizing agent	bleaching agent	A	eExam

<input type="checkbox"/>	MCQ	Flourine has all but one of these properties	The oxidation state of is -1	It can oxidise all the other halide ions to their respective elements	The size is larger than the other halogens	it is the strongest oxidising agent in the whole group	C	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	Which of these statements is not true of halogens ?	They are only one electron short of the noble gas configuration	They form the cation X ⁺	Fluorine is the strongest oxidising agent in the whole group	The bond in interhalogen compounds is polar	B	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	Which of these is used in making crucible?	graphite	Diamond	coke	coal	A	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	Which of these compounds is used as antiknock additives to petrol for internal combustion engines?	lead sulphate	lead oxide	Tetramethyl lead	lead chloride	C	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	The property by which elements form long chain compounds by single or multiple bond formation between atoms of the same elements is----	Polymorphysm	Polimerization	Catenation	Dention	C	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	Which of these is not a component of lead-acid battery?	trehydrooxsulphate(VI) aci	lead oxide	lead chloride	spongy lead	C	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	Which of these elements of group IV can form compounds easily in +2 oxidation state	C	Si	Pb	Ge	C	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	Activated charcoal is used	in cutting	as decolourizing agent	in jewellery	in refrigeration	B	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	High purity silicon is obtained from-----	Silicon hydroxide	Silicon tetrachloride	Silicon dihydride	Silicon dioxide	B	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	Which of these is not a form of carbon?	Coke	charcoal	carbon black	carbon hydride	D	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	Artificial graphite can be produced on a large scale by heating	Coke with silica	Carbon with silica	Charcoal with silica	Carbon with silica	A	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	Artificial diamond can be produced by quenching----	a solution of carbon in lead	a solution of silicon in lead	a solution of carbon in tin	a solution of carbon in iron	D	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	Which of these elements shows polymorphism?	Si	C	Sn	Pb	C	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	An example of amorphous form of carbon is----	Carbon monoxide	Diamond	Lampblack	Graphite	C	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	Each carbon in Diamond is bonded to	6 Carbon atoms	4 Carbon atoms	3 Carbon atoms	2 Carbon atoms	B	<input type="checkbox"/> eExam
<input type="checkbox"/>	MCQ	Diamond behaves as an insulator because-----	It is a solid	It is made of carbon	It does not have mobile electrons	It is used as an jewelry	C	<input type="checkbox"/> eExam

<input type="checkbox"/>								
<input type="checkbox"/>	MCQ	The three dimensional linkages makes diamond one of the	hardest	most beautiful	most expensive	slimmest	A	eExam
<input type="checkbox"/>	MCQ	The structure of graphite is- ----	Irregular planar hexagonal	Regular planar hexagonal	Irregular tetrahedral octagonal	Regular tetrahedral planar	B	eExam
<input type="checkbox"/>	MCQ	Which of these two elements are important components of semiconductors and transistors?	,Ge and Sn	Sn and Pb	Si and Pb	Si and Ge	D	eExam
<input type="checkbox"/>	MCQ	The two allotropies of carbon are-----	Charcoal and Coke	Coke and Diamond	Diamond and Charcoal	Diamond and Graphite	D	eExam
<input type="checkbox"/>	MCQ	Clays are essentially alumina-silicates of ----	Na and K	Ca and Si	Na or Ca	Al or Si	C	eExam
<input type="checkbox"/>	MCQ	Allotropy is the existence of elements in more than-----	One form	One physical state	One oxide compound	One hydride compound	A	eExam

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