

eExam Question Bank

Coursecode:

Choose Coursecode

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<input type="checkbox"/>	Question Type	Question	A	B	C	D	Answer	Remark
<input type="checkbox"/>	FBQ	The label on a stock bottle of an acid reads: 56% by mass and 1.25 specific gravity. If the molar mass of the acid is 56, the volume of this acid that is required to prepare 250cm^3 of 1.0 molar concentration of the acid is <input type="text"/>	2.0 centimetre cube	2cm^3				<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	If actual yield of an ester is 32.7g and theoretical yield is 35.1g .The percentage yield of the ester will be <input type="text"/>	0.932					<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	The number of gram-equivalent weight of solute in one cubic decimeter of solution is <input type="text"/>	Normality					<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	Calculate the formular mass of potassium carbonate.(Relative atomic masses : K=39.10amu, C=12.01amu, 16.00amu). <input type="text"/>	138.21amu					<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	6.02×10^{23} is called <input type="text"/>	Avogadro number					<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	Calculate the percentage by mass of Oxygen in CH_3OH <input type="text"/>	0.4995					<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	The amount of energy required to break a particular bond is called <input type="text"/>	Bond dissociation energy					<input type="button" value="eExam"/>

<input type="checkbox"/>								
<input type="checkbox"/>	FBQ	In a chemical reaction, the balanced chemical equation which shows the quantitative relationship between masses of reactants and products is known as <input type="text"/> .-	Stoichiometry of the reaction					eExam
<input type="checkbox"/>	FBQ	The amount of product expected from given amounts of starting materials is called <input type="text"/> .-	Theoretical yield					eExam
<input type="checkbox"/>	FBQ	The amount of product obtained from an experiment is called <input type="text"/> .-	actual yield					eExam
<input type="checkbox"/>	FBQ	Calculate the percentage by mass of Hydrogen in NH_3 <input type="text"/> .-	0.1775					eExam
<input type="checkbox"/>	FBQ	In a solution, the component in smaller amount is called <input type="text"/> .-	solute					eExam
<input type="checkbox"/>	FBQ	In balancing this equation $\text{Na}_2\text{CO}_3 \rightarrow ? \text{NaOH} + \text{CaCO}_3$. ? is <input type="text"/> .-	2					eExam
<input type="checkbox"/>	FBQ	The formula for caustic soda is <input type="text"/> .-	NaOH					eExam
<input type="checkbox"/>	FBQ	Mixtures with particle sizes greater than 1000 nanometres are called <input type="text"/> .-	Suspensions					eExam
<input type="checkbox"/>	FBQ	The label on a stock bottle of acid reads : 56% by mass and 1.25 specific gravity. If the molar mass of the acid is 56, what is the the concentration in grams per dm^3 <input type="text"/> .-	700g					eExam
<input type="checkbox"/>	FBQ	Mixtures with particle sizes between 2 to 1000 nanometre are called <input type="text"/> .-	colloids					eExam
<input type="checkbox"/>	FBQ	Convert 500g of Na_2CO_3 to moles <input type="text"/> .-	4.717mol	4.72mol				eExam
<input type="checkbox"/>	FBQ	When propane reacts with oxygen, the products are <input type="text"/> and <input type="text"/> .-	carbon dioxide, water.	CO_2 , H_2O				eExam
<input type="checkbox"/>	FBQ	The molar mass of Na_2CO_3 <input type="text"/> .-	106g					eExam

<input type="checkbox"/>								
<input type="checkbox"/>	FBQ	Convert 2 moles of NaOH to grams of NaOH. Answer is <input type="text"/> .	80g					eExam
<input type="checkbox"/>	FBQ	Dissolution Of sodium chloride in water results in <input type="text"/> type of mixture.	Homogeneous					eExam
<input type="checkbox"/>	FBQ	A mixture in which the components are evenly distributed and the composition is uniform throughout the mixture is called <input type="text"/> mixture.	Homogeneous					eExam
<input type="checkbox"/>	FBQ	When two or more elements combine chemically in fixed proportion by mass, <input type="text"/> is formed.	Compounds					eExam
<input type="checkbox"/>	FBQ	Calculate the $[H^+]$ of an aqueous solution with $[OH^-]$ of $1 \times 10^{-10} M$. What is the pH of the solution? Is the solution acidic or basic? <input type="text"/>	4					eExam
<input type="checkbox"/>	FBQ	An acid HA has a pka of 4.5. What is the concentration of H_3^+ in 0.110M solution of the acid ? <input type="text"/>	1.876×10^{-3}					eExam
<input type="checkbox"/>	FBQ	An acid HA has a pka of 4.5, what is the Ka value? <input type="text"/>	3.2×10^{-5}					eExam
<input type="checkbox"/>	FBQ	Calculate the $[OH^-]$ of an aqueous solution of Ph=11 <input type="text"/>	$1 \times 10^{-3} M$					eExam
<input type="checkbox"/>	FBQ	The number of protons or electrons in the atom of an element is equal to the <input type="text"/> of the element.	Atomic number					eExam
<input type="checkbox"/>	FBQ	The actual yield of product (in g or moles) expressed as a percentage of the theoretical yield (in g or moles) is called <input type="text"/> .	Percentage yield					eExam
<input type="checkbox"/>	FBQ	A high percentage yield implies that <input type="text"/> .	Reaction was successful					eExam
<input type="checkbox"/>	FBQ	The reagent that will not make any significant contribution to the theoretical yield is <input type="text"/> .	Limiting reagent					eExam

<input type="checkbox"/>								
<input type="checkbox"/>	FBQ	A separation technique that is used to isolate a desired solid from a solid -liquid mixture or for freeing a desired liquid of solid impurities is called <input type="text"/>	Filtration					eExam
<input type="checkbox"/>	FBQ	The process by which traces of water is removed by treating the liquid with suitable drying agents is called <input type="text"/>	Dehydration					eExam
<input type="checkbox"/>	FBQ	An apparatus that can be used to remove traces of water from a substance is <input type="text"/>	Dessicator					eExam
<input type="checkbox"/>	FBQ	The condensate collected during the vaporization of liquid is called <input type="text"/>	Distillate					eExam
<input type="checkbox"/>	FBQ	The separation technique employed to separate mixtures of two or more liquids with slightly different boiling points is called <input type="text"/>	Fractional distillation					eExam
<input type="checkbox"/>	FBQ	The technique employed for the separation of colours is called <input type="text"/>	Chromatography					eExam
<input type="checkbox"/>	FBQ	The technique of separation employed to purify an organic solid that may be contaminated by impurities is called <input type="text"/>	Recrystallization					eExam
<input type="checkbox"/>	FBQ	The base unit of a measured liquid called volume is <input type="text"/>	Cubic meter					eExam
<input type="checkbox"/>	FBQ	An undergraduate weighed out 20grams of sodium hydroxide pellets. If Na =23, O = 16 andH = 1, What is the mole of this sodium hydroxide. <input type="text"/>	0.25mole					eExam
<input type="checkbox"/>	FBQ	<input type="text"/> is the quantitative isolation of a substance by precipitation and weighing of the precipitate.	Gravimetric analysis					eExam
<input type="checkbox"/>	FBQ	Gravimetric analysis can be generalized into <input type="text"/> types.	two					eExam

<input type="checkbox"/>								
<input type="checkbox"/>	FBQ	The weakly bound water in efflorescent compound is known as <input type="text"/> ..	water of crystallization					eExam
<input type="checkbox"/>	FBQ	Pure substances combine with water in a fixed mole ratio to yield compounds called <input type="text"/> ..	Hydrates					eExam
<input type="checkbox"/>	FBQ	<input type="text"/> is the amount of heat released per amount of fuel.	Heat of combustion					eExam
<input type="checkbox"/>	FBQ	<input type="text"/> represent a large portion of the world energy supply.	Fossil fuels					eExam
<input type="checkbox"/>	FBQ	<input type="text"/> substances are hydrates that lose water when simply exposed to the atmosphere.	Efflorescent					eExam
<input type="checkbox"/>	FBQ	<input type="text"/> substances readily absorb moisture from the air and are used as drying agents.	Hygroscopic					eExam
<input type="checkbox"/>	FBQ	The complex ion formed above is <input type="text"/> in colour.	reddish					eExam
<input type="checkbox"/>	FBQ	The complex ion formed by Aluminium and hydroxyl ion is known as <input type="text"/>	alu -oxy ion					eExam
<input type="checkbox"/>	FBQ	The atoms in a compound are held together by <input type="text"/> ..	chemical bonds					eExam
<input type="checkbox"/>	FBQ	<input type="text"/> is the simplest way to express information about the atoms that constitute any given chemical compound.	empirical formular					eExam
<input type="checkbox"/>	FBQ	Reactions i which substances undergo changes in oxidation number are referred to as <input type="text"/> ..	redox reaction	Oxidation-reduction reaction				eExam
<input type="checkbox"/>	FBQ	<input type="text"/> is the drying agent used only at temperatures below 30 °C	Sodium sulphate					eExam
<input type="checkbox"/>	FBQ	<input type="text"/> is the most effective drying agent.	calcium chloride					eExam
<input type="checkbox"/>	FBQ	If an acid is splashed on your skin, wash at once with <input type="text"/> ..	plenty of water					eExam

<input type="checkbox"/>								
<input type="checkbox"/>	FBQ	Distillation is used to remove [] from a solution.	solvent					eExam
<input type="checkbox"/>	FBQ	Ionic compounds that crystallize from solutions containing sulfate ion, a trivalent cation and a monovalent cation are known as [] .	Alums					eExam
<input type="checkbox"/>	FBQ	The complex ion formed by Aluminium and hydroxyl ion is known as [] .	Aluminate					eExam
<input type="checkbox"/>	MCQ	What is the oxidation state of iodine in iodate ion?	\$\$+5\$\$	\$\$+6\$\$	\$\$+7\$\$	\$\$+8\$\$	C	eExam
<input type="checkbox"/>	MCQ	When diluting acids, always pour _____.	water into acids	acids into water	hydrogen into acids	acids into oxygen	B	eExam
<input type="checkbox"/>	MCQ	_____ must not be carried through a group of students.	Hot equipment or dangerous chemicals	used beakers	unwashed containers	neutral results	A	eExam
<input type="checkbox"/>	MCQ	Students with open skin wounds on hands must _____.	rub oil	wear gloves	wash wounds	none of these	B	eExam
<input type="checkbox"/>	MCQ	If any equipment is not working properly, you must _____.	repair it	inform the instructor	inform your colleague	work with it	B	eExam
<input type="checkbox"/>	MCQ	Each laboratory activity should be Prepared by reading _____.	all books	few instructions	all instructions	few books	C	eExam
<input type="checkbox"/>	MCQ	_____ is used whenever a reaction mixture has to be kept boiling for an appreciable time and the solvent is volatile	reflux condenser	bottled condenser	glass condenser	covered condenser	A	eExam
<input type="checkbox"/>	MCQ	The stoichiometric ratio is _____.	the mole ratio of the products	the mole ratio of the catalysts	the mole ratio of the reactants	none of these	C	eExam
<input type="checkbox"/>	MCQ	A _____ is the reagent that is completely consumed during a chemical reaction.	catalytic reactant	limiting reactant	spontaneous reactant	all of these	B	eExam
<input type="checkbox"/>	MCQ	The percentage of nitrogen in atomsphere is _____.	0.5	0.78	0.21	0.17	B	eExam
<input type="checkbox"/>	MCQ	To determine the percent composition, the amounts of each gas can be measured both by _____.	weight and volume	mass and volume	weight and density	volume and density	A	eExam
<input type="checkbox"/>	MCQ	A student measuring a gas should be careful of these except _____.	Atmospheric pressure.	water temperatures	molecular movement	laboratory temperatures	C	eExam
<input type="checkbox"/>	MCQ	The total mass of each element in a compound depends on the number of _____.	its atoms	its electrons	its proton	its neutron	A	eExam

<input type="checkbox"/>								
<input type="checkbox"/>	MCQ	_____ helps identify each constituent element by its chemical symbol and indicates the number of atoms of each element in the compound.	molecular formular	chemical formular	nuclear arrangement	Bonding	B	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	_____ is a homogenous mixture of gases such as nitrogen, oxygen, argon, and trace amounts of other elemental gases and carbon dioxide.	Fumes	Vapours	Air	water	C	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	The atoms in a compound are held together by _____.	chemical reactions	atomic arrangement	nuclear bond	chemical bonds	D	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	_____ is the simplest way to express information about the atoms that constitute any given chemical compound.	empirical formular	molecular formular	chemical formular	functional formular	A	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	_____ may be used to remove any brown stains left on the glassware.	nitric acid	Oxalic acid	sulphuric acid	organic acids	A	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	An algebraic decrease in oxidation number or a process in which electrons are gained _____.	Oxidation	Reduction	Redox	none of these	B	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	An algebraic increase in oxidation number, or a process in which electrons are lost _____.	oxidation	reduction	redox	none of these	A	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	Reactions in which substances undergo changes in oxidation number are referred to as _____.	reduction reaction	oxidation reaction	biased reaction	redox reaction	D	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	_____ is the drying agent used only at temperatures below 30 °C.	calcium oxide	Sodium sulphate	potassium sulphate	calcium chloride	B	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	_____ is the most effective drying agent.	calcium chloride	potassium chloride	calcium oxide	sodium chloride	A	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	A vacuum desiccator is used to _____.	collect solid samples	speed up the dissolution of samples	separate two liquids	speed the drying of a sample	D	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	In setting up of apparatus, which of these is not a way of loosening seized joints.	soak the joint in water, then try tapping	rock the cone in the socket	tap the joint gently with a block of wood	warm the joint in a small flame, then tap gently	A	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	If a solid to be weighed has big crystals or lumps, we are to _____.	force it in to the weighing bottle	compress it with an iron	pound it in a mortar	soak it in water	C	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	_____ is the amount of heat released per amount of fuel .	Heat of vapourization	Heat of combustion	Heat of solution	Heat of reaction	B	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	_____ substances are hydrates that lose water when simply exposed to the atmosphere.	Deliquescent	Hygroscopic	Efflorescent	Illuminscent	A	<input type="button" value="eExam"/>

<input type="checkbox"/>								
<input type="checkbox"/>	MCQ	_____ substances readily absorb moisture from the air and are used as drying agents.	Deliquescent	Hygroscopic	Efflorescent	Illumiscent	B	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	_____ substances absorb water from the air until they form a solution.	Deliquescent	Hygroscopic	Efflorescent	Illumiscent	A	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	_____ is an alkaline substance, and reacts with conc. HCl to form Boric acid.	Bromine	Borax	Butane	Bromide	B	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	The product formed between a cation and some multidentate ligand to form a ring structure is called a _____.	filterate	product	chelate	precipitate	C	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	_____ are formed by reactions between cations and anions.	Salts	Compounds	Acids	Bases	A	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	_____ precipitating agent is used to precipitate sulphate ion.	sodium chloride	barium chloride	magnesium nitrate	silver nitrate	B	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	_____ precipitating agent is used to precipitate halide ions such as chloride.	sodium chloride	barium chloride	magnesium nitrate	silver nitrate	D	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	The empirical formula does not indicate the exact number of _____ in a single molecule.	ions	protons	electrons	atoms	D	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	There are _____ types of hardness of water.	2	3	4	5	A	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	_____ is the indicator that enables the detection of when the EDTA has completely chelated the metal impurities.	Methyl orange	Phenophthalein	Eriochrome Black T	All of these	C	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	_____ is a substance whose molecules can form several bonds to a single metal ion.	A chelating agent	A precipitating agent	A dehydrating agent	A drying agent	A	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	_____ ions make the most significant contribution to water hardness.	Phosphorous	Calcium	Potassium	Oxygen	B	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	_____ is the substance to be analysed.	A chelate	An anhydride	An analyte	A sample	C	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	Gravimetric analysis by definition is based upon the measurement of _____.	depth	length	weight	mass	D	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	_____ are crystalline salts that are bonded to water molecules in definite proportions.	Hydrides	Amphoterics	Anhydrides	Hydrates	D	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	_____ is used for the removal of coloured impurities from crude materials.	Recrystallization	activated coal	activated charcoal	Distillation	C	<input type="button" value="eExam"/>

<input type="checkbox"/>								
<input type="checkbox"/>	MCQ	The energy rating of a fossil fuel gives us information about its _____.	empirical formular	molecular composition	chemical formular	chemical composition	D	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	A very important property of a fuel is its _____.	heat of reaction	heat of vapourization	heat of combustion	heat of solution	C	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	_____ is a light black residue consisting of carbon, and any remaining ash, obtained by removing water and other volatile constituents from animal and vegetation	Coal	Charcoal	Coke	Graphite	B	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	_____ is the loss of water (or a solvent) of crystallization from a hydrated or solvated salt to the atmosphere on exposure to air.	Hydroscopy	Efflorescence	Deliquescence	Illuminscence	B	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	_____ the process in which a soluble substance picks up water vapor from the air to form a solution.	Hydroscopy	Efflorescence	Deliquescence	Illuminscence	C	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	_____ is the ability of a substance to attract and hold water molecules from the surrounding environment.	Hydroscopy	Efflorescence	Deliquescence	Illuminscence	A	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	Borax is treated with conc. HCl because it is _____ in nature.	Volatile	insoluble	Soluble	inflammable	A	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	The major requirements for a good gravimetric analysis are these except _____.	the reagent will react only with the analyte of interest to form a precipitate	it forms two and only two product with the analyte	that the analyte precipitates quantitatively from solution, that is,	it forms one and only one product with the analyte	B	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	Organic precipitating reagents contain functional groups that combine with inorganic ions to form _____.	basic salts	acidic salts	soluble salts	insoluble salts	D	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	Potassium, ammonium, rubidium, and cesium ions can be precipitated by _____.	silver nitrate	sodium chloride	barium chloride	sodium tetraphenylborate	D	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	Precipitation is effected by _____ types precipitating agents.	3	4	2	1	C	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	_____ analysis is concerned with the determination of a substance by the process of weighing.	potentiometric	Complexometric	Volumetric	Gravimetric	D	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	Hard water contains one of these metal ions.	$[\text{Ca}^{2+}]$	$[\text{Al}^{3+}]$	$[\text{K}^{+}]$	$[\text{O}^{2-}]$	A	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	Temporary hardness of water is removed by _____.	Distillation	Boiling	Filtration	Crystallization	B	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	To establish the exact amounts of various metal ions in a sample of hard water, a _____ is required.	volumetric titration	complexometric titration	Conductometric titration	potentiometric titration	B	<input type="button" value="eExam"/>

<input type="checkbox"/>								
<input type="checkbox"/>	MCQ	Potassium aluminum sulphate dodecahydrate is also called_____.	Potash	Alum	Edible salt	Industrial sugar	B	eExam

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