

eExam Question Bank

Coursecode:

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<input type="checkbox"/>	Question Type	Question	A	B	C	D	Answer	Remark
<input type="checkbox"/>	FBQ	In chemistry, natural or synthetic substances that change colour in response to the nature of chemical environment are called <input type="text"/>	Indicators					<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	All phenols give a colour when ferric chloride solution is added. <input type="text"/> (True or False)						<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	In BENEDICT TEST (for aldehydes and sugars reducing sugars) give <input type="text"/> precipitate when heated with benedict's solution	brick- red					<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	In the BEILSTEIN TEST for halogenated compounds, halogens give a transient <input type="text"/> colour	green					<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	2,4-DINITROPHENYLHYDRAZINE TEST (for aldehydes and ketones) will be positive for an aldehyde or ketone as indicated by the formation of <input type="text"/> precipitate	yellow					<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	To classify an organic compound into a given family requires first detecting a specific <input type="text"/> in the molecules of organic compounds	functional group					<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	Dehydration of an alcohol is a common method of introducing <input type="text"/> into an organic compound.	unsaturation					<input type="button" value="eExam"/>

<input type="checkbox"/>									
<input type="checkbox"/>	FBQ	Potassium permanganate, a purple solution loses colour with alkene and changes to <input type="text"/>	brown						eExam
<input type="checkbox"/>	FBQ	Cyclohexanol is a volatile acid <input type="text"/> (True or False)							eExam
<input type="checkbox"/>	FBQ	<input type="text"/> is defined as the time taken for a component to go from injection to detection	Retention time						eExam
<input type="checkbox"/>	FBQ	Two types of detector used in gas Chromatography are: thermal conductivity detectors and <input type="text"/>	flame ionisation detection						eExam
<input type="checkbox"/>	FBQ	Molecules with different polarity partition to different extents <input type="text"/> (true or false)	1						eExam
<input type="checkbox"/>	FBQ	In column chromatography, the non polar compounds move very slowly <input type="text"/> (True or false)							eExam
<input type="checkbox"/>	FBQ	In Column Chromatography, the stationary phase is packed into a glass tube to form a <input type="text"/> of granules	column						eExam
<input type="checkbox"/>	FBQ	An organic solvent or a mixture of solvents known as <input type="text"/> flows down through the column	the eluent						eExam
<input type="checkbox"/>	FBQ	In TLC, the surface consists of a very thin layer of <input type="text"/> on plastic or aluminium backing	silica gel						eExam
<input type="checkbox"/>	FBQ	In paper chromatography <input type="text"/> serves as a support for immobile liquid phase	filter paper						eExam
<input type="checkbox"/>	FBQ	Separation of two sample components in chromatography is based on their different distribution between two <input type="text"/>	non-miscible phases						eExam
<input type="checkbox"/>	FBQ	In liquid chromatography, the mobile phase is <input type="text"/>	Liquid						eExam

<input type="checkbox"/>									
<input type="checkbox"/>	FBQ	Tollens' reagent is used to detect the presence of <input type="text"/>	aldehydes						eExam
<input type="checkbox"/>	FBQ	NaOH solution is highly corrosive acid <input type="text"/> (true or false)							eExam
<input type="checkbox"/>	FBQ	Sulfuric acid is used as a catalyst for the esterification reactions <input type="text"/> (True or False)	1						eExam
<input type="checkbox"/>	FBQ	All the organic compounds used or produced in this experiment are highly <input type="text"/> (True or False)							eExam
<input type="checkbox"/>	FBQ	<input type="text"/> give fruits and flowers subtle odour/aroma	Esters						eExam
<input type="checkbox"/>	FBQ	What does STP stand for <input type="text"/>	Standard Temperature and Pressure						eExam
<input type="checkbox"/>	FBQ	In the reaction of magnesium with hydrochloric acid <input type="text"/> gas is generated with Magnesium Chloride	Hydrogen						eExam
<input type="checkbox"/>	FBQ	The value of universal gas constant (R) will differ depending on the units used for pressure and <input type="text"/>	volume						eExam
<input type="checkbox"/>	FBQ	The ideal gas equation of state is also known as <input type="text"/>	ideal gas law						eExam
<input type="checkbox"/>	FBQ	In paper chromatography, water forms hydrogen bonds with the fibers of the paper <input type="text"/> (true or false)							eExam
<input type="checkbox"/>	FBQ	In paper chromatography, water serves as the mobile phase, <input type="text"/> (True or False)							eExam
<input type="checkbox"/>	FBQ	Paper chromatography is a form of partition chromatography . <input type="text"/> _(True or False)	1						eExam

<input type="checkbox"/>									
<input type="checkbox"/>	FBQ	In paper chromatography, the more strongly a substance is adsorbed, the closer it will stay near the origin. <input type="text"/> (True or False)	1						eExam
<input type="checkbox"/>	FBQ	TLC is useful because it is reproducible. <input type="text"/> _(True or False)	1						eExam
<input type="checkbox"/>	FBQ	Amino acids are colourless compounds <input type="text"/> (True or False)	1						eExam
<input type="checkbox"/>	FBQ	In alcohol dehydration reactions, according to <input type="text"/> rule the major product is usually the more highly substituted alkene	Zaitsev's Rule						eExam
<input type="checkbox"/>	FBQ	When heated with strong acids catalysts (most commonly H ₂ SO ₄ , H ₃ PO ₄), alcohols typically undergo a 1,2-elimination reactions to generate an alkene and <input type="text"/>	water						eExam
<input type="checkbox"/>	FBQ	Cyclohexanol can be dehydrated to <input type="text"/>	cyclohexene						eExam
<input type="checkbox"/>	FBQ	Dehydration of an alcohol can follow either the E2 or <input type="text"/> _the mechanism	E1						eExam
<input type="checkbox"/>	FBQ	Four basic types of chemical reactions in organic chemistry: combination, elimination, substitution, and <input type="text"/>	rearrangement						eExam
<input type="checkbox"/>	FBQ	In thin layer chromatography, a solid phase, is also called <input type="text"/>	the adsorbent						eExam
<input type="checkbox"/>	FBQ	Chromatographic separations take advantage of the fact that different substances are partitioned differently between two phases, a mobile phase and a <input type="text"/>	stationary phase						eExam
<input type="checkbox"/>	FBQ	<input type="text"/> is a heated injector port, where the sample is vaporised if necessary	Injection system						eExam

<input type="checkbox"/>								
<input type="checkbox"/>	FBQ	<input type="text"/> of aldehydes yields carboxylic acids	Oxidation					eExam
<input type="checkbox"/>	FBQ	<input type="text"/> are easily oxidized a fact due to the presence of the hydrogen attached to the carbonyl group	Aldehydes					eExam
<input type="checkbox"/>	FBQ	<input type="text"/> is a commercial solvent and is used in paint thinners and nail polish removers	Acetone					eExam
<input type="checkbox"/>	FBQ	In gas chromatography the mobile phase <input type="text"/>	gas					eExam
<input type="checkbox"/>	FBQ	Another name for acetaldehyde is <input type="text"/>	ethanal					eExam
<input type="checkbox"/>	FBQ	<input type="text"/> may be prepared by oxidation of 1° alcohols	Aldehydes					eExam
<input type="checkbox"/>	FBQ	In the preparation of esters, the temperature should be maintained at about <input type="text"/>	70	70°C				eExam
<input type="checkbox"/>	FBQ	The ester found in orange is <input type="text"/>	octyl acetate					eExam
<input type="checkbox"/>	FBQ	Aluminum metal melts at 660.37°C. What is the temperature in Kelvin? <input type="text"/>	933.52	933.52K				eExam
<input type="checkbox"/>	FBQ	From the Law of Conservation of energy, during any physical or chemical change: Energy lost = <input type="text"/>	Energy gained					eExam
<input type="checkbox"/>	FBQ	Heat is evolved in a reaction if the reaction is <input type="text"/>	exothermic					eExam
<input type="checkbox"/>	FBQ	Convert 300K to Celsius <input type="text"/>	26.85	26.85°C				eExam
<input type="checkbox"/>	FBQ	Convert 100K to Fahrenheit <input type="text"/>	-279.67	-279.67F				eExam
<input type="checkbox"/>	FBQ	Vinegar will turn methyl red <input type="text"/>	red					eExam
<input type="checkbox"/>	FBQ	Phenol Red turns <input type="text"/> with acids	yellow					eExam

<input type="checkbox"/>									
<input type="checkbox"/>	FBQ	Phenolphthalein turns <input type="text"/> with acid	colourless						eExam
<input type="checkbox"/>	FBQ	When a nonvolatile solute is dissolved in the liquid, the vapour pressure of the liquid is <input type="text"/>	decreased						eExam
<input type="checkbox"/>	FBQ	<input type="text"/> is an instrument used to measure heat flow in and out of a system.	Calorimeter						eExam
<input type="checkbox"/>	MCQ	One of these reagents can be used to distinguish I II III alcohols_____	Jones reagent	Lucas' reagent	Tollen's reagent	Benedicts reagent	B		eExam
<input type="checkbox"/>	MCQ	The SI unit for heat is _____	calorie	celsuis	centigrade	joule	D		eExam
<input type="checkbox"/>	MCQ	The energy change of a reaction that occurs under a constant pressure is defined as the _____	entropy change	enthalpy change	adiabatic change	conservational change	B		eExam
<input type="checkbox"/>	MCQ	Which of these statements is not true ? _____	The freezing point of a liquid is depressed when it contains a dissolved solid	salt water will freeze lower than the normal freezing point of water	The addition of ethylene glycol to the water in a car radiator lowers its boiling point	the molar mass of an unknown solute can be determined using freezing-point depression	C		eExam
<input type="checkbox"/>	MCQ	A solution with a low pH value is called _____	an alkali	a base	an acid	a caustic	C		eExam
<input type="checkbox"/>	MCQ	lemon juice will turn methyl red _____	red	blue	colourless	yellow	A		eExam
<input type="checkbox"/>	MCQ	Natural or synthetic substances that change colour in response to the nature of chemical environment are known as _____	indicators	monitors	catalysts	adsorbents	A		eExam
<input type="checkbox"/>	MCQ	Ammonia is a base and thus will turn methyl red_____	blue	yellow	red	orange	B		eExam
<input type="checkbox"/>	MCQ	Thymol Blue gives a _____ colour with bases	red	blue	yellow	orange	C		eExam
<input type="checkbox"/>	MCQ	Methyl orange gives a _____ colour with acids	Red	orange	yellow	colourless	A		eExam
<input type="checkbox"/>	MCQ	Which of these tests is used to distinguish primary, secondary and tertiary alcohols of six carbons or less ? _____	iodoform test	lucas test	hinsberg test	tollens test	B		eExam
<input type="checkbox"/>	MCQ	In the preparation of cyclohexene,tetraoxo sulphate(IV) is used as a _____	acidifying agent	catalyst	drying agent	saturating agent	B		eExam

<input type="checkbox"/>	MCQ	One of these is used as adrying agent in the preparation of cyclohexene _____	calcuim chloride	cyclohexanol	hydrochloric acid	tetraoxo sulphate(IV)	A	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	All but one of these reagents can be reduced by aldehydes only _____	Fehling's solution	Tollen's reagent	Jones reagent	Brady's reagent	D	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	In the reaction of halogens with carbon tetrachloride a violet colour indicates the presence of _____	iodine	chlorine	bromine	flourine	A	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	Which of these statements is not true of compounds in the S1 solubility group _____	These are very polar compounds	These are neutral compounds	They consist of salts of carboxylic acids	salts of amines belong to this group	B	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	In an ignition test ,sooty flame means _____	compound is solid	compound is aliphatic	compound is aromatic	compound is electrovalent	C	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	Functional groups influence all but one of these properties organic compounds _____	physical	chemical	spectral	production date	D	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	In the elimination of water from an alcohol _____ are formed	ketones	aldehydes	alkenes	alkanes	C	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	In the seperation of amino acids using paper chromatography ,ninhydrin gives _____ colour the amino acids	green	purple	yellow	brown	B	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	In paper chromatography, filter paper serves as a support for _____	immobile liquid phase	mobile liquid phase	immobile gas phase	mobile solid phase	A	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	What is the work of Calcium Chloride in the preparation of cyclohexene using cyclohexanol _____	to maintain a white coloration	remove traces of water	maintain the temperature below 100o	to make less soluble in water	B	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	One of these apparatus is required in the production of cyclohexene using cyclohexanol _____	distillation set up	Erlenmyer flask	Soxlet apparatus	thermometer adaptor	C	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	In dehydration of cyclohexanol to produce cyclohexene all but one of these reagents are required _____	sodium chloride	potassuim permanganate	copper sulphate	phosphoric acid	C	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	The dehydration of alcohols to give alkenes is an example of _____	elimination	substitution	rearrangement	combination	A	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	All but one of these are basic reactions of organic chemistry _____	elimination	substitution	rearrangement	titration	D	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	In TLC, the glass can be coated by all but one of the following _____	silica gel	sodium hydroxide	aluminum oxide	cellulose	B	<input type="button" value="eExam"/>

<input type="checkbox"/>	MCQ	In TLC, the ratio of the distance the compound travels to the distance the solvent travels is called _____	Rf value	Df value	R value	Q value	A	eExam
<input type="checkbox"/>	MCQ	In TLC the stationary phase is _____	liquid absorbent	solid absorbent	gas absorbent	colloid	B	eExam
<input type="checkbox"/>	MCQ	_____ can be used to analyse samples taken from athletes to check for the presence of drugs	paper chromatography	thin layer chromatography	gas chromatography	HPLC	C	eExam
<input type="checkbox"/>	MCQ	All but one of these are components of the HPLC. _____	Solvent Reservoir	Carrier gas	The Pump System	The Detector	B	eExam
<input type="checkbox"/>	MCQ	Retention time is proportional to the _____	the temperature of the compound	colour of the compound	size of the compound	viscosity of the compound	C	eExam
<input type="checkbox"/>	MCQ	Retention time depends on all but one of the following _____	The nature of and the interactions between the solute and the stationary and mobile phases	diameter of the column	The temperature of the column	The temperature of the carrier gas	D	eExam
<input type="checkbox"/>	MCQ	_____ is a type of detector used in gas chromatography	heat ionisation detection	flame ionisation detection	gas conductivity detector	flame conductivity detectors	B	eExam
<input type="checkbox"/>	MCQ	One of these is not a basic component of the Gas chromatograph _____	Pneumatic system	Oven	Column	Gas	D	eExam
<input type="checkbox"/>	MCQ	Silica gel can be used in one of these types of chromatography _____	Gas	HPLC	column	paper	C	eExam
<input type="checkbox"/>	MCQ	Oxidation of aldehydes yields _____	ketones	alcohols	carboxylic acids	alkalis	C	eExam
<input type="checkbox"/>	MCQ	In paper chromatography, the solvent pulled up through the paper by one of these _____	atomic bond	magnetic force	capillary action	stationary phase	C	eExam
<input type="checkbox"/>	MCQ	In liquid chromatography the mobile phase _____	gas	water	liquid	cellulose fibre	C	eExam
<input type="checkbox"/>	MCQ	Separation of two sample components in chromatography is based _____	similarity with each other	their different distribution between two non-miscible phases	solubility in the mobile phase	solubility in the stationary phase	B	eExam
<input type="checkbox"/>	MCQ	The ester found in orange is _____	methyl butyrate	isobutyl	octyl acetate	benzyl acetate	C	eExam
<input type="checkbox"/>	MCQ	During preparation of Acetone from 2-Propanol the temperature is maintained at _____	100o C	80o C	50o C	30 oC	C	eExam

<input type="checkbox"/>	MCQ	In paper chromatography, the stationary phase is _____	filter paper	liquid that is held on the fibers of the paper	liquid in developing solvent	cellulose	B	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	Which of these is not a type of chromatography _____	gas	high performance liquid	thin layer	bilayer	D	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	All but one of these are materials necessary for the laboratory preparation of ethanal _____	Safety glasses	Tongs	Bunsen burner	water bath	D	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	Michaelis-Bauer is associated with one of the following techniques _____	chromatography	Distillation	electrophoresis	sublimation	A	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	The differential distribution of the components of the mixture between a stationary phase and a mobile phase can be achieved by the method of _____	Distillation	chromatography	electrophoresis	sublimation	B	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	In the reaction, $\text{CH}_3\text{-COOH} + \text{HO-CH}_2\text{CH}_3 \leftrightarrow \text{CH}_3\text{-COO-CH}_2\text{CH}_3 + \text{H}_2\text{O}$ the product is _____	an aldehyde	a ketone	an ester	an acid	B	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	In the preparation of acetone all but one of these is a requirement _____	isopropyl alcohol	acidic dichromate	Distilled water	Sodium Hydroxide	D	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	Which of these statements is not true	Tollens' reagent is used to detect the presence of aldehydes	Oxidation of aldehydes yields carboxylic acids	Benedict's reagent can oxidize aldehydes	aldehydes and ketones can form hydrogen bonds one to another	D	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	Another name for acetaldehyde is _____	ethanal	acetone	ethanol	Acetic acid	A	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	In the laboratory preparation of acetone one of these is necessary _____	copper sulphate	ethanol	2-propanol	acetic acid	C	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	Ketones may be prepared by _____ of 2° alcohols.	oxidation	reduction	dehydration	esterification	A	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	_____ may be prepared by oxidation of 1° alcohols	Esters	Aldehydes	Ketones	Ethers	B	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	In the preparation of esters, the temperature should be maintained at about _____	70 °C	150 °C	100 °C	30 °C	A	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	Ethyl butyrate is an ester found in _____	apple	orange	banana	pineapple	D	<input type="button" value="eExam"/>
<input type="checkbox"/>	MCQ	One of these is a catalyst in used esterification reactions _____	sulphuric acid	Methanoic acid	acetic acid	water	A	<input type="button" value="eExam"/>

<input type="checkbox"/>								
<input type="checkbox"/>	MCQ	Which of these is not an apparatus required for the preparation of ester_____	water bath	hotplate	pipette	sulphuric acid	D	eExam
<input type="checkbox"/>	MCQ	Water bath is used in the preparation of esters because_____	Water boils faster in a water bath	Most of the reactants and products are highly flammable	More ester is produced with water bath	Ester have strong odour	B	eExam
<input type="checkbox"/>	MCQ	The chief organic product when an organic acid is heated with an alcohol in the presence of a strong mineral acid is a member of the family of organic compounds known as _____	ether	ester	alcohol	soap	B	eExam

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