

eExam Question Bank

Coursecode:

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<input type="checkbox"/>	Question Type	Question	A	B	C	D	Answer	Remark
<input type="checkbox"/>	FBQ	The <input type="text"/> of an element is the the mass of one atom of the element compared to (1/12) of the mass of one atom of carbon - 12 isotopes	relative atomic mass					<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	Chemical symbols and formulae are abbreviations used to represent <input type="text"/>	elements and compounds					<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	<input type="text"/> of a compound is the sum of the relative atomic masses of the elements present in the chemical formular of the compound	relative molecular mass					<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	With the use of the <input type="text"/> it has been possible to determine fairly accurately the relative atomic masses of elements	mass spectrometer					<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	The <input type="text"/> of an element is the the mass of one atom of the element compared to (1/12) of the mass of one atom of carbon - 12 isotopes	relative atomic mass					<input type="button" value="eExam"/>
<input type="checkbox"/>	FBQ	For the atom to be electrically neutral, the numbers of <input type="text"/> must be equal	protons and electrons					<input type="button" value="eExam"/>

<input type="checkbox"/>									
<input type="checkbox"/>	FBQ	The constituents of the atom are <input type="text"/> , <input type="text"/> , and <input type="text"/>	protons, electrons, neutrons						eExam
<input type="checkbox"/>	FBQ	The Proton carries <input type="text"/> magnitude of charge as the electron	the same						eExam
<input type="checkbox"/>	FBQ	The proton is the <input type="text"/> charged particle	positively						eExam
<input type="checkbox"/>	FBQ	The <input type="text"/> charged particle in matter is the electron	negatively						eExam
<input type="checkbox"/>	FBQ	Evidences about the nature of matter came from results of experiments of <input type="text"/> , <input type="text"/> , and <input type="text"/>	Faraday, Thompson, Millikan						eExam
<input type="checkbox"/>	FBQ	Matter is <input type="text"/> in nature	electrical						eExam
<input type="checkbox"/>	FBQ	The postulation that in a chemical change atoms are neither created nor destroyed refers to <input type="text"/> __ Theory	Dalton's Atomic						eExam
<input type="checkbox"/>	FBQ	Matter is <input type="text"/> in nature	electrical						eExam
<input type="checkbox"/>	FBQ	The existence of atoms of the same element having different masses finds relevance with Atoms of the same element having different masses are called <input type="text"/>	isotopes						eExam

<input type="checkbox"/>									
<input type="checkbox"/>	FBQ	The statement, matter is composed of more fundamental particles, some of which are electrically neutral, some carry positive charge and some negative charge relates to <input type="text"/>	Daltons Atomic theory						eExam
<input type="checkbox"/>	FBQ	<input type="text"/> law states that when atoms combine they do so in simple ratios	Boyles law						eExam
<input type="checkbox"/>	FBQ	<input type="text"/> law states that atoms can neither be created nor destroyed	Daltons Atomic theory						eExam
<input type="checkbox"/>	FBQ	<input type="text"/> law states that all elements are made up of small, indivisible particles called atoms	Daltons Atomic theory						eExam
<input type="checkbox"/>	FBQ	A <input type="text"/> is a tested proposal to explain an observed statement of facts	theory						eExam
<input type="checkbox"/>	FBQ	The wave mechanical treatment of the atom overcomes the limitation of <input type="text"/> models	Niel's Bohr						eExam
<input type="checkbox"/>	FBQ	Allowed transitions are transitions from one orbit to another and will lead to the emmission/absorption of <input type="text"/>	energy						eExam
<input type="checkbox"/>	FBQ	<input type="text"/> proposed a model for the atom in which electrons move round the nucleus only in allowed orbits	Niel's Bohr						eExam
<input type="checkbox"/>	FBQ	<input type="text"/> proposed that the atom consisted of a tiny positively charged nucleus	Rutherford						eExam

<input type="checkbox"/>									
<input type="checkbox"/>	FBQ	The law of <input type="text"/> states that when two different compounds are formed from the same two elements, the masses of one element which react with a fixed mass of the other are in a ratio of small whole numbers	multiple propotion						eExam
<input type="checkbox"/>	FBQ	The implication of the law of <input type="text"/> a chemical equation must always be balanced to account for all atoms present on the reactant side, and on the product side of the reaction.	conservation of matter						eExam
<input type="checkbox"/>	FBQ	The law of <input type="text"/> states that matter is neither created nor destroyed in a chemical reaction	conservation of matter						eExam
<input type="checkbox"/>	FBQ	The law of <input type="text"/> States that the proportion by mass of the different elements present is fixed for a pure sample of the compound irrespective of the method of preparations	constant composition						eExam
<input type="checkbox"/>	FBQ	The <input type="text"/> of a compound gives the proportion of the different elements present in the compound by mass	formula						eExam
<input type="checkbox"/>	FBQ	Chemical symbols and formulae are abbreviations used to represent <input type="text"/>	the first one or two letters of the name of the element						eExam
<input type="checkbox"/>	FBQ	Chemical equations are a <input type="text"/>	summary of a chemical reaction						eExam
<input type="checkbox"/>	FBQ	<input type="text"/> exist even between uncombined atoms and non polar molecules	Covalent bonding						eExam

<input type="checkbox"/>									
<input type="checkbox"/>	FBQ	A form of bonding which arises from the sharing of electrons among atoms, resulting in a form of bond <input type="text"/>	Covalent bonding						eExam
<input type="checkbox"/>	FBQ	The statement that atoms tend to gain or lose electrons until there are eight electrons in their valence shell refers to <input type="text"/>	Octet rule						eExam
<input type="checkbox"/>	FBQ	<input type="text"/> involves electron transfer from the valence shell of one atom to the valence shell of the other	Electrovalent bonding						eExam
<input type="checkbox"/>	FBQ	Sterilising of hospital equipments is a use for <input type="text"/>	radioactivity						eExam
<input type="checkbox"/>	FBQ	<input type="text"/> Is a process in which the nucleus of a heavy element is spilt into two nuclei of nearly equal mass with a release of energy and radiation.	nuclear fission						eExam
<input type="checkbox"/>	FBQ	<input type="text"/> a process in which two or more light nuclei combine to form a heavier nucleus with a release of energy	nuclear fussion						eExam
<input type="checkbox"/>	FBQ	Gamma rays are deflected and have <input type="text"/>	no effect on electric field						eExam
<input type="checkbox"/>	FBQ	<input type="text"/> rays are deflected towards a -ve pole	Alpha						eExam
<input type="checkbox"/>	FBQ	<input type="text"/> rays are deflected towards a +ve pole	Beta						eExam
<input type="checkbox"/>	FBQ	<input type="text"/> rays have very low ionising power	gamma						eExam
<input type="checkbox"/>	FBQ	<input type="text"/> rays have low ionising power	Beta						eExam

<input type="checkbox"/>								
<input type="checkbox"/>	FBQ	<input type="text"/> rays are negatively charged	Beta					eExam
<input type="checkbox"/>	FBQ	<input type="text"/> rays are neutrally charged	gamma					eExam
<input type="checkbox"/>	FBQ	<input type="text"/> rays have high ionising power	Alpha					eExam
<input type="checkbox"/>	FBQ	Alpha rays are <input type="text"/> Charged	positively					eExam
<input type="checkbox"/>	FBQ	A radioactive element in its decay emits alpha, beta and <input type="text"/>	gamma rays					eExam
<input type="checkbox"/>	FBQ	The half-life is a measure of the <input type="text"/>	time taken for half of the radioactive substance to decay					eExam
<input type="checkbox"/>	FBQ	Unstable nuclei account for <input type="text"/>	radioactivity					eExam
<input type="checkbox"/>	FBQ	The knowledge of radioactivity was instrumental to the debunking of <input type="text"/>	Dalton's postulate	Dalton's theory				eExam
<input type="checkbox"/>	MCQ	What is the basic unit of matter in chemical reactions ?	compound	atom	solid	molecule	B	eExam
<input type="checkbox"/>	MCQ	How would you explain chemical reactions of atoms	combinations and rearrangement	involving reactions	chemical union	non of the advanced options	A	eExam
<input type="checkbox"/>	MCQ	What is the smallest unit of a compound that has the characteristics of the compound	compound	atom	solid	molecule	D	eExam
<input type="checkbox"/>	MCQ	What is the smallest unit of an element that can take part in a chemical reaction.	compound	atom	solid	molecule	B	eExam
<input type="checkbox"/>	MCQ	What is the smallest unit of an element that has the properties and characteristics of the element ?	compound	atom	solid	molecule	B	eExam
<input type="checkbox"/>	MCQ	Which of these methods can be used to separate a mixture	dissolution	cooling	filtration	all of the listed	D	eExam

<input type="checkbox"/>									
<input type="checkbox"/>	MCQ	How would you separate a mixture with varying components?	chemical action only	physical methods	orthodox methods	none of the listed approaches	B		eExam
<input type="checkbox"/>	MCQ	What is defined by a physical combination of elements or compounds.	A mixture	A compound	An element	A solid	A		eExam
<input type="checkbox"/>	MCQ	What is the number of compounds available.	specific	limited	reserved	limitless	D		eExam
<input type="checkbox"/>	MCQ	How much energy is often required to split compounds into the constituent elements	minimal	no amount	a lot of	commensurate amount of	C		eExam
<input type="checkbox"/>	MCQ	Which of these statements is true about the properties of compounds and the elements from which they are formed.	same with	different from	similar to	not related to	B		eExam
<input type="checkbox"/>	MCQ	What results when two or more elements combine chemically in fixed proportion by mass.	compound	mixture	element	gases	A		eExam
<input type="checkbox"/>	MCQ	Non metals	cannot conduct heat	cannot conduct electricity	can be either solids or liquids	all these characteristics are applicable	D		eExam
<input type="checkbox"/>	MCQ	Non metals are	solids only	solids and gases	gases only	liquids only	B		eExam
<input type="checkbox"/>	MCQ	Identify the characteristics of a non metal	have characteristic lustre	lack lustre	are good conductors	all of these are applicable to non metals	B		eExam
<input type="checkbox"/>	MCQ	At what temperature is metal solid	High temperature	At all times	None of the listed conditions	At room temperature	D		eExam
<input type="checkbox"/>	MCQ	Which of these is a general characteristic of metals	Lack lustre	Lustre	Fluid	Non conductor	B		eExam
<input type="checkbox"/>	MCQ	How many groups can you subdivide an element into ?	Four groups	Three groups	Two groups	Five groups	C		eExam
<input type="checkbox"/>	MCQ	What is a pure substance that can be broken down into elements called?	Element	Mixture	Compound	Atoms	C		eExam
<input type="checkbox"/>	MCQ	Which of these statements hold true for an element ?	cannot be split	can be split	is divisible	none of these	A		eExam
<input type="checkbox"/>	MCQ	What do chemical symbols and formulae used to represent ?	atom	molecules	elements and compounds	compounds only	C		eExam

<input type="checkbox"/>									
<input type="checkbox"/>	MCQ	The relative molecular mass of a compound is the	sum of the atomic masses of the elements therein	sum of the atomic masses of the elements present in the chemical formula of the compound	sum of the relative atomic masses of the elements present in the chemical formula of the compound	sum of the atomic masses of the molecules present in the chemical formula of the compound	C		eExam
<input type="checkbox"/>	MCQ	What instrument would you use to determine fairly accurately the relative atomic masses of elements	spectrophotometer	mass spectrometer	electronic balances	laboratory beam balance	B		eExam
<input type="checkbox"/>	MCQ	What is the relative atomic mass of an element?	average atomic mass of two or more atoms	the equal mass of the atom against carbon 12	the mass of one atom of the element compared to (1/12) of the mass of one atom of carbon - 12 isotopes	average of the masses of the protons, neutrons and electrons	C		eExam
<input type="checkbox"/>	MCQ	For the atom to be electrically neutral, what should be the magnitude of the charges of the protons and electrons?	number of electrons should be equal to those of the protons and neutrons	number of protons should be equal to those of the electrons and neutrons	number of protons and neutrons must be equal	number of protons and electrons must be equal	D		eExam
<input type="checkbox"/>	MCQ	What are the constituents of the atom?	electron	neutron	proton	protons, electrons and neutrons	D		eExam
<input type="checkbox"/>	MCQ	What is the third particle is the neutron, a neutral particle	electron	nucleus	neutron	proton	C		eExam
<input type="checkbox"/>	MCQ	What is the magnitude of the charge on a proton and an electron?	the same	a lesser	different	an incomparable	A		eExam
<input type="checkbox"/>	MCQ	What is the nature of the charge on a proton?	negatively	neutrally	positively	heavily	C		eExam
<input type="checkbox"/>	MCQ	What is the nature of the charges on an electron?	negatively	neutrally	positively	heavily	A		eExam
<input type="checkbox"/>	MCQ	Which of these scientists conducted experiments that accounted for the evidences about the nature of matter?	Michael faraday	Thompson	Millikan	Faraday, Thompson and Millikan	D		eExam
<input type="checkbox"/>	MCQ	Which of these is the nature of Matter?	electrical	chemical	atomic	electrical, chemical and atomic	A		eExam

<input type="checkbox"/>									
<input type="checkbox"/>	MCQ	Which law does the postulation that in a chemical change atoms are neither created nor destroyed defines?	Charles law	Gay Lusaacs law	Boyles law	Daltons Atomic theory	D		eExam
<input type="checkbox"/>	MCQ	What are atoms of the same element having different masses known as?	Molecule	mixtures	compounds	isotopes	D		eExam
<input type="checkbox"/>	MCQ	The existence of atoms of the same element having different masses finds relevance with	Daltons Atomic theory	Charles law	Gay Lusaacs law	Boyles law	A		eExam
<input type="checkbox"/>	MCQ	The statement, matter is composed of more fundamental particles, some of which are electrically neutral, some carry positive charge and some negative charge relates to	Daltons Atomic theory	Charles law	Gay Lusaacs law	Boyles law	A		eExam
<input type="checkbox"/>	MCQ	Which law states that when atoms combine they do so in simple ratios	Daltons Atomic theory	Charles law	Gay Lusaacs law	Boyles law	A		eExam
<input type="checkbox"/>	MCQ	Which law states that atoms can neither be created nor destroyed	Charles law	Boyles law	Daltons Atomic theory	Gay Lusaacs law	B		eExam
<input type="checkbox"/>	MCQ	Which law states that all elements are made up of small, indivisible particles called atoms	Charles law	Daltons Atomic theory	Boyles law	Gay Lusaacs law	B		eExam
<input type="checkbox"/>	MCQ	What is a tested proposal to explain an observed statement of facts	law	hypothesis	theory	conjecture	C		eExam
<input type="checkbox"/>	MCQ	What form of bonding exist even between uncombined atoms and non polar molecules?	Covalent bonding	Dative bonding	Electrovalent bonding	Vaander Waal's forces	D		eExam
<input type="checkbox"/>	MCQ	What form of bonding arises from the sharing of electrons among atoms?	Covalent bonding	Dative bonding	Electrovalent bonding	Vaander Waal's forces	A		eExam
<input type="checkbox"/>	MCQ	What does the statement that atoms tend to gain or lose electrons until there are eight electrons in their valence shell refers to?	law of conservation of matter	Hunds multiplicity rule	Octet rule	law of definate proportion	C		eExam
<input type="checkbox"/>	MCQ	What form of bonding involves electron transfer from the valence shell of one atom to the valence shell of the other	Covalent bonding	Dative bonding	Electrovalent bonding	Chemical bonding	C		eExam

<input type="checkbox"/>									
<input type="checkbox"/>	MCQ	Which of these is a use for radioactivity?	treatment of cancer	sterilising hospital equipments	tracing movement of a substance in a process	all of the above are applicable	D		eExam
<input type="checkbox"/>	MCQ	The process in which the nucleus of a heavy element is spilt into two nuclei of nearly equal mass with a release of energy and radiation is known as	nuclear fission	nuclear fussion	atomisation	ionisation	A		eExam
<input type="checkbox"/>	MCQ	The process in which two or more light nuclei combine to form a heavier nucleus with a release of energy is known as	nuclear fussion	nuclear fission	atomisation	ionisation	A		eExam
<input type="checkbox"/>	MCQ	Identify the path on which Gamma rays are deflected	towards a +ve pole	towards a -ve pole	have no effect on electric field	ricochetted back in its path	C		eExam
<input type="checkbox"/>	MCQ	Identify the path of deflection of Alpha rays	deflected towards a +ve pole	deflected towards a -ve pole	no effect on electric field	ricochetted back in its path	B		eExam
<input type="checkbox"/>	MCQ	Identify the path of deflection of Beta rays	deflected towards a +ve pole	deflected towards a -ve pole	no effect on electric field	ricochetted back in its path	A		eExam
<input type="checkbox"/>	MCQ	Which of these statements hold true about Gamma rays?	high ionising power on gases	low ionising power on gases	very low ionising power on gases	both negatively and positively charged	C		eExam
<input type="checkbox"/>	MCQ	Which of these statements hold true about Beta rays?	high ionising power on gases	low ionising power on gases	very low ionising power on gases	both negatively and positively charged	B		eExam
<input type="checkbox"/>	MCQ	What type of charges do Beta rays carry?	negatively charged	positively charged	neutrally charged	both negatively and positively charged	A		eExam
<input type="checkbox"/>	MCQ	What is the nature of the charge on Gamma rays?	negatively charged	positively charged	neutrally charged	both negatively and positively charged	C		eExam
<input type="checkbox"/>	MCQ	Identify the correct statement about Alpha rays	high ionising power on gases	low ionising power on gases	very low ionising power on gases	no ionising power on gases	A		eExam
<input type="checkbox"/>	MCQ	What type of charges do Alpha rays bear?	negatively charged	positively charged	neutrally charged	both negatively and positively charged	B		eExam

<input type="checkbox"/>								
<input type="checkbox"/>	MCQ	what radiations does a radioactive element in its decay emit?	Alpha rays only	Beta rays only	Gamma rays only	Alpha, beta and gamma rays	D	eExam
<input type="checkbox"/>	MCQ	Which of these statements best define the phenomenon of half-life?	measure of the instability of an electron	a measure of the life span of an isotope	a measure of the time taken for half of the radioactive substance to decay	a measure of the life expectancy of an atom	C	eExam
<input type="checkbox"/>	MCQ	Which of these can be associated with unstable nuclei?	chemical reactions	radioisotopes	radioactivity	photo emissions	C	eExam
<input type="checkbox"/>	MCQ	Which of these laws was discarded as a result of The knowledge of radioactivity?	Dalton's postulate	John Dewey's assertion	Gay Lussac's law	Charles law	A	eExam

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