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**NATIONAL OPEN UNIVERSITY OF NIGERIA**

**14-16 AHMADU BELLO WAY, VICTORIA ISLAND LAGOS**

**MARCH/APRIL 2016 EXAMINATION**

 **SCHOOL OF SCIENCE AND TECHNOLOGY**

**COURSE CODE: CHM391**

**COURSE TITLE: PRACTICAL CHEMISTRY V – INORGANIC AND ANALYTICAL**

**TIME: 2 HOURS**

**INSTRUCTION: ANSWER ANY FOUR QUESTIONS**

**QUESTION ONE**

1. Use a schematic diagram to show the procedure for the separation and identification of a metal cation in a given sample by qualitative inorganic analysis of cation method.

 101/2 marks

1. Differentiate between qualitative analytical groups and groups of the periodic table.

 7 marks

 **QUESTION TWO**

1. Below is an infrared spectrum of an unknown organic compound. Using the table of the characteristic infrared absorption bands of organic functional groups provided below, identify the functional groups present in this organic compounds.



 **Table : CHARACTERISTIC INFRARED ABSORPTION BANDS OF FUNCTIONAL GROUPS**

|  |  |  |  |
| --- | --- | --- | --- |
| Class of compounds | Absorption cm-1 | Intensity | Assignment |
| Alkanes and AkylsAlkenesAlkynesEthersAldehydesCarboxylic acidsAlcohols | 2850 – 30001450 – 14701370 – 13901365 +1395Two bands715 - 7253020-31401640-1670910+990Two bands885-895665-730960-980790-8403265-33352100-2140610-7002190-22601085-11501020-1075 and1200-1275(Two bands)2700-27251720-17401685-17102500-35001710-17151680-17103300 – 34001125 - 1205 | ssmmww-mvw-mm+ssm-s, broadsssmsvw-wsmmsss bends, broads, broads, broadm - s | C – H stretchC- H bendCH2, C – H bend-CH(CH3)3 bend-(CH2)n bend=C-H StretchC=C Stretch=C-H bend=C-H bend=C-H bend=C-H bend=C-H bend=C-H StretchC=C Stretch=C-H bendC=C StretchC-O-C Stretch=C-O-C sym and asym stretchH-C=O StretchC=O StretchC=O StretchO-H StretchC=O StretchC=O StretchO – H stretchC –O stretch |

**Intensity abbreviations**: vw = very weak, w = weak, m = medium, s = strong, vs = very strong

 9 marks

1. Explain briefly the principle of infrared Spectroscopy

 81/2marks

**QUESTION THREE**

1. A laboratory analysis was conducted to determine the amount of aspirin (2-acetylbenzoic acid) in a commercial aspirin product using UV-Visible Spectroscopic method. The commercial aspirin was complexed with sodium hydroxide and iron (iii) Solution to form the salicylate dianion (the intensity of the colour of the complex is directly related to the concentration of aspirin present) and the absorbance of this complex determined. Five series of solutions with different aspirin concentrations was prepared and labeled A-E and complexed. The absorbance of each solution was measured. Using the information below
2. Construct a calibration curve
3. Use the calibration curve to determine the amount of aspirin in the commercial aspirin product

|  |  |  |
| --- | --- | --- |
| Concentrations of standard solution (mg/ml) | Absorbances of standard solution | Standard solutions |
| 0.05 | 0.2 | A |
| 0.10 | 0.4 | B |
| 0.15 | 0.6 | C |
| 0.20 | 0.8 | D |
| 0.25 | 1.0 | E |

The absorbance of the commercial aspirin product is 0.5

 171/2 marks

**QUESTION FOUR**

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1. Write briefly on the principle of Atomic Absorption Spectroscopy

 8 marks

1. The concentration of magnesium in tap water is to be determined by employing Atomic Absorption Spectroscopy (AAS) to measure the absorbance of magnesium in tap water. A set of standard solutions of magnesium was prepared and their absorbance obtained. Prepare a calibration curve of the magnesium standard solutions and determine from the calibration curve the concentration of magnesium in the tap water. The absorbance of the magnesium in the tap water read 0. 2

|  |  |
| --- | --- |
| Concentration of magnesium standard solutions (PPM) | Absorbance of magnesium standard solutions |
| 2.50 | 0.2 |
| 4.500 | 0.3 |
| 6.500 | 0.5 |
| 8.500 | 0.7 |
| 10.500 | 0.9 |

 91/2 marks

 **QUESTION FIVE**

1. In an analysis to determine the chloride present in a given sample weighing 1.52g by precipitation gravimetric method, aqueous solution of the sample was acidified with dilute acid and a slight excess of silver nitrate solution was added, whereupon the chloride present in the sample was precipited as silver chloride. If the weight of the silver chloride precipitate obtained is 0.126 g, calculate the percentage of chloride in the sample. Gravimetric factor = Cl/Agcl =0.24737

 9 marks

1. What is digestion of a precipitate and why is it necessary

 81/2 marks

**QUESTION SIX**

1. The concentration of an organic compound is to be determined from its calibration curve by UV-visible spectroscopy, if the wavelength of maximum absorption (λ max) of this compound is unknown, determine the wavelength of maximum (λ max) at which the calibration curve can be prepared using the information provided below.

|  |  |
| --- | --- |
| Absorbance(s) of the organic compound | Wavelengths of absorption of the organic compound (nm)  |
| 0.100 | 360 |
| 0.110 | 380 |
| 0.120 | 400 |
| 0.125 | 420 |
| 0.130 | 440 |
| 0.16 | 460 |
| 0.165 | 480 |
| 0.400 | 500 |
| 0.60 | 520 |
| 1.00 | 540 |
| 1.10 | 560 |
| 0.80 | 580 |
| 0.40 | 600 |
| 0.10 | 620 |
| 0.11 | 640 |
| 0.12 | 660 |

 10marks

1. Describe briefly the principle of UV-visible spectroscopy

 71/2 marks.