

**NATIONAL OPEN UNIVERSITY OF NIGERIA**

**UNIVERSITY VILLAGE, PLOT 91 CADASTRAL ZONE, NNAMDI AZIKIWE EXPRESS WAY,**

 **JABI - ABUJA.**

 **FACULTY OF SCIENCES**

 **DEPARTMENT OF PURE AND APPLIED SCIENCE**

 **JULY 2018 EXAMINATIONS**

**COURSE CODE: CHM 391**

**COURSE TITLE: PRACTICAL CHEMISTRY V – INORGANIC AND ANALYTICAL**

**COURSE UNIT: 2**

**TIME: 2 HOURS**

INSTRUCTION: **Question one is compulsory. Answer question one and any other**

 **three questions.**

**QUESTION ONE**

1ai) Write briefly on the principle of Atomic Absorption Spectroscopy. 6 marks

1aii) List the kind of compounds that are capable of absorbing UV-Visible radiation. 3marks

1aiii) Among the compounds listed in (1aii) above, which of them strongly absorb UV-Visible radiation or absorb UV-Visible radiation the most. 1mark

1bi) Below is a table of values obtained for the determination of total alkalinity in a water sample, use this table to calculate the total alkalinity of a 100 ml water sample titrated with 0.02 M H2SO4.

|  |  |  |  |  |
| --- | --- | --- | --- | --- |
| NO of titration  | Volume of sample (mL)  | Initial burette reading  | Final burette reading  | Volume of Sulphuric (mL)  |
| 1  | 100  | 0.00  | 8.50  | 8.50  |
| 2  | 100  | 0.00  | 8.30  | 8.30  |
| 3  | 100  | 0.00  | 8.30  | 8.30  |

 5 marks

1bii) Mention the indicators used to detect the presence of the following sources of alkalinity in water samples.

1. OH-
2. HCO3-  2 marks

1biii) Give two significance and application of alkalinity. 2 marks

1ci) Outline the experimental procedure of how to determine the functional groups present in an organic sample. 5 marks

1cii) What kind of molecules absorb infrared radiation. 1 mark

**QUESTION TWO**

2a) In an experiment to determine the concentration of lead in soil collected from the side of a road using atomic absorption spectrophotometer, standard lead solutions were prepared and the absorbance of each solution measured. The data below was obtained

|  |  |  |
| --- | --- | --- |
| Standard solutions | Concentration (ppm) | Absorbance |
| Blank | 0.00 | 0.00 |
| Standard 1 | 1.00 | 0.17 |
| Standard 2 | 2.00 | 0.34 |
| Standard 3 | 3.00 | 0.48 |
| Standard 4 | 4.00 | 0.65 |
| Standard 5 | 5.00 | 0.83 |

If the absorbance of the road side soil sample was 0.58, determine the concentration of lead in the soil sample from a calibration curve prepared by you. 131/2 marks

2b) Mention one use of atomic absorption spectroscopy (AAS). 11/2 marks

**QUESTION THREE**

In an experiment to determine the acidity of a water sample, the burette was filled with 0.02M NaOH, 100 ml of the water sample was transferred to a conical flask using pipette and few drops of methyl orange indicator was added to the content in the conical flask. The sample was titrated against the 0.02M NaOH solution until the orange colour became faint. The volume V1 consumed for this titration was noted. To the same solution in the conical flask few drops of phenolphthalein indicator was added and the titration continued until the colour changed to faint pink colour. The total volume V2 consumed for this titration was noted. The experiment was repeated three times to get concordant values.

The values obtained are presented in table 1 and 2. Use these values to calculate:

1. The Mineral Acidity of the water sample
2. Total Acidity of the water sample

Table 1 for Mineral Acidity

|  |  |  |  |  |
| --- | --- | --- | --- | --- |
| NO of titration  | Volume of sample (mL)  | Initial burette reading  | Final burette reading  | Volume of NaOH (mL)  |
| 1  | 100  | 0.00  | 0.50  | 0.50  |
| 2  | 100  | 0.00  | 0.40  | 0.40  |
| 3  | 100  | 0.00  | 0.40  | 0.40  |

Table 2 for Total Acidity

|  |  |  |  |  |
| --- | --- | --- | --- | --- |
| NO of titration  | Volume of sample (mL)  | Initial burette reading  | Final burette reading  | Volume of NaOH (mL)  |
| 1  | 100  | 0.00  | 2.20  | 2.20  |
| 2  | 100  | 0.00  | 2.30  | 2.30  |
| 3  | 100  | 0.00  | 2.30  | 2.30  |

The molecular weight of CaCO3 = 100g (=100,000mg) 15 marks

**QUESTION FOUR**

4a) Identify the functional groups present in the infrared spectrum given below using the table of characteristic infrared absorption bands of organic functional groups provided below.



**CHARACTERISTIC INFRARED ABSORPTION BANDS OF FUNCTIONAL GROUPS**

|  |  |  |  |
| --- | --- | --- | --- |
| Class of compounds | Absorption cm-1 | Intensity | Assignment |
| Alkanes and AkylsAlkenesAlkynesEthersAldehydesCarboxylic acidsAlcohols | 2850 – 30001450 – 14701370 – 13901365 +1395Two bands715 - 7253020-31401640-1670910+990Two bands885-895665-730960-980790-8403265-33352100-2140610-7002190-22601085-11501020-1075 and1200-1275(Two bands)2700-27251720-17401685-171011002500-35001710-17151680-17103300 – 34001125 - 1205 | ssmmww-mvw-mm+ssm-s, broadsssmsvw-wsmmssms bends, broads, broads, broadm - s | C – H stretchC- H bendCH2, C – H bend-CH(CH3)3 bend-(CH2)n bend=C-H StretchC=C Stretch=C-H bend=C-H bend=C-H bend=C-H bend=C-H bend=C-H StretchC=C Stretch=C-H bendC=C StretchC-O-C Stretch=C-O-C sym and asym stretchH-C=O StretchC=O StretchC=O StretchC-C-C bendingO-H StretchC=O StretchC=O StretchO – H stretchC –O stretch |

**Intensity abbreviations**: vw = very weak, w = weak, m = medium, s = strong, vs = very strong

 12 marks

4b) In what way is potentiometric titration different from classical titration. 3 marks

 **QUESTION FIVE**

5a) The concentration of an organic compound is to be determined using UV-visible spectrophotometer, if the wavelength of maximum absorption (λ max) of this organic compound is unknown, use the information provided below to obtain the wavelength of maximum (λ max) at which the determination will be carried out.

|  |  |
| --- | --- |
| Absorbance(s) of the organic compound | Wavelengths of absorption of the organic compound (nm)  |
| 0.100 | 360 |
| 0.110 | 380 |
| 0.120 | 400 |
| 0.125 | 420 |
| 0.130 | 440 |
| 0.160 | 460 |
| 0.165 | 480 |
| 0.400 | 500 |
| 0.60 | 520 |
| 1.00 | 540 |
| 1.10 | 560 |
| 0.80 | 580 |
| 0.40 | 600 |
| 0.10 | 620 |
| 0.11 | 640 |
| 0.12 | 660 |

 12 marks

5b) In what way is potentiometric titration different from classical titration. 3 marks